

1. Determine if a pure substance is made of atoms or molecules.

_____ C	_____ H ₂ O	_____ carbon dioxide	_____ sodium
_____ Ar	_____ Cu	_____ neon	_____ sulfur
_____ PBr ₃	_____ P ₄	_____ sulfur trioxide	_____ neon
_____ Ca	_____ N ₂	_____ silicon	_____ fluorine
_____ H ₂	_____ C ₆ H ₁₂ O ₆	_____ boron trioxide	_____ iron

2. Distinguish between pure substances and mixtures.

_____ NaCl & H ₂ O	_____ sand & sugar
_____ CO ₂ & O ₂	_____ gold
_____ CH ₄	_____ lead (II) nitrate dissolved in water
_____ 2Fe(C ₂ H ₃ O ₂) ₂ · 10H ₂ O	_____ nitrogen trihydride and oxygen
_____ Ca ²⁺ & SO ₄ ²⁻ & H ₂ O	_____ marbles and beads

3. Distinguish between elements and compounds.

_____ I ₂	_____ nitrogen monoxide
_____ Ag ₂ P	_____ nickel
_____ Sr ₃ N ₂	_____ chromium (III) sulfide
_____ CO	_____ aluminum bromide
_____ Cu	_____ fluorine

4. Distinguish between homogeneous mixtures (solutions) and heterogeneous mixtures.

- i. How can you tell if a mixture is homogeneous or heterogeneous?

5. Describe a chemical on the macroscopic, particle, and symbolic levels.

- i. Know how to represent a chemical with particle pictures.
ii. Know how to represent a chemical with symbols (formulas).

6. Know which elements are diatomic in nature and which are polyatomic in nature.

- i. Know the name and symbol of each diatomic (7) and polyatomic element (2) in nature.

7. Know the names and symbols of the first 54 elements on the periodic table.

- i. Be familiar with the first 54 elements so you can easily locate it on the periodic table.
ii. If you have them memorized, you will save time having to look it up on the periodic table.

8. Understand how to use the periodic table to identify groups/families of elements (and their chemical properties) known as alkali metals, alkaline earth metals, halogens, and noble gases.

- i. Which groups are alkali metals, alkaline earth metals, halogens, and noble gases found?

9. Distinguish between metals and nonmetals.

_____ H	_____ Cl	_____ lithium	_____ silver
_____ N	_____ Cr	_____ oxygen	_____ carbon
_____ Sr	_____ P	_____ arsenic	_____ sulfur

10. Understand the nature of the covalent bond and ionic bond.

- i. Covalent bonds are formed when two nonmetals fight over or share electrons.
ii. Ionic bonds are formed when a cation (being positively charged because it lost electrons) is attracted to an anion (being negatively charged because it gained electrons).
a. Ionic formulas are written in simplest whole number ratios.

Honors Chemistry Study Guide

RS Peters



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Barron's Science 360: A Complete Study Guide to Chemistry with Online Practice Barron's Educational Series, Mark Kernion, Joseph A. Mascetta, 2021-09-07 Barron's Science 360 Chemistry is your complete go to guide for everything chemistry This comprehensive guide is an essential resource for High school and college courses Homeschooling Virtual Learning Learning pods Inside you'll find Comprehensive Content Review Begin your study with the basic building block of chemistry and build as you go Topics include atomic structure chemical formulas electrochemistry the basics of organic chemistry and much more Effective Organization Topic organization and simple lesson formats break down the subject matter into manageable learning modules that help guide a successful study plan customized to your needs Clear Examples and Illustrations Easy to follow explanations hundreds of helpful illustrations and numerous step by step examples make this book ideal for self study and rapid learning Practice Exercises Each chapter ends with practice exercises designed to reinforce and extend key skills and concepts These checkup exercises along with the answers and solutions will help you assess your understanding and monitor your progress Access to Online Practice Take your learning online for 50 practice questions designed to test your knowledge with automated scoring to show you how far you have come **ACT Study Guide Premium, 2025: 6 Practice Tests + Comprehensive Review + Online Practice** Barron's Educational Series, Brian Stewart, 2025-02-04 Get ready for ACT test day with Barron's and crush your goals Barron's ACT is the most up to date and comprehensive guide available to students who want to showcase their college readiness earn top scholarships and gain admission to the most competitive universities Internationally known expert author and tutor Brian W Stewart a Princeton graduate and perfect ACT score holder puts his 30 000 plus hours of teaching and tutoring experience to work for you He gives you the same clear and concise advice to excel on the ACT that has helped his students from all ability levels earn perfect ACT scores and admission to Ivy League universities This fully updated guide includes over 2 000 practice questions and a wide ranging review of ACT subject material to target your weak areas and enhance your strengths 4 full length practice tests including a diagnostic test with a self assessment to target specific question types for your customized study 2 additional full length practice tests online for further practice Detailed overview of the ACT with comprehensive answers to frequently asked questions and detailed advice for students who have extended time accommodations Study plan recommendations based on the amount of time you have to prepare Review of all the concepts tested on the ACT and in depth grammar instruction including punctuation parallelism and wordiness Advanced drills to practice the toughest types of problems you will face on test day Proven strategies to help you with time management minimizing careless mistakes avoiding overthinking and why determining to read or not to read is the essential for a successful approach to ACT science passage The 4 C method for answering ACT reading questions and how to adjust your ACT reading technique for literary narratives and informational passages ACT Writing strategies with plenty of sample

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,1998 *High School Chemistry Teachers Magazine* ,1967

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