

Honors Chemistry Worksheet 3

Stoichiometry Practice Problems

Name _____ Period _____ Date _____

Instructions: Balance the following chemical equations and then determine the missing information for each of the conditions given. The four questions related to each equation are independent of one another. Answers for a particular numbered problem should be in the same units as the information given (i.e. grams to grams, moles to moles, particles to particles).



- | | | | | | | |
|----|-------------------|--|-------------------|--|---------------------------------------|--|
| 1. | <u>5mol</u> | | <u> </u> | | <u> </u> | |
| 2. | <u> </u> | | <u>2.5mol</u> | | <u> </u> | |
| 3. | <u>25g</u> | | <u> </u> | | <u> </u> | |
| 4. | <u> </u> | | <u> </u> | | <u>2.41x10²⁴ molecules</u> | |



- | | | | | | | |
|----|-------------------|--|-------------------|--|-------------------|--|
| 1. | <u> </u> | | <u>.75mol</u> | | <u> </u> | <u> </u> |
| 2. | <u>50g</u> | | <u> </u> | | <u> </u> | <u> </u> |
| 3. | <u> </u> | | <u> </u> | | <u>3.5mol</u> | <u> </u> |
| 4. | <u> </u> | | <u> </u> | | <u> </u> | <u>1.204x10²³ molecules</u> |



- | | | | | | | |
|----|-----------------------------------|--|-------------------|--|-------------------|--|
| 1. | <u>1.51x10²³ atoms</u> | | <u> </u> | | <u> </u> | |
| 2. | <u> </u> | | <u>1.25mol</u> | | <u> </u> | |
| 3. | <u> </u> | | <u> </u> | | <u>25 grams</u> | |

Honors Chemistry Stoichiometry Test Answers

Doctor Chemistry



Honors Chemistry Stoichiometry Test Answers:

A Whole Year of Chemistry Quizzes Siegfried Howell, 2019-12-12 This book *A Whole Year of Chemistry Quizzes* was written to provide easy to use and grade quizzes to assess the comprehension of honors students Advance Placement students AP and International Baccalaureate IB students This book of quizzes guides the teacher and the student through what is required in a non watered down chemistry course that leads students towards test and college readiness The outline of this book has a minimum of 4 quizzes per chapter that prepares students for the formative assessment associated at the end of all chapters The 25 chapters include topics that are covered in the honors chemistry setting as well as specialty topics like thermodynamics kinetics rates of reactions that are seen in the Advance Placement classes Included within this book are quizzes for the International Baccalaureate teacher that wishes to test students on environmental chemistry as well as biological and food chemistry This is a book that was written to fill the void of valuable resources needed for novice and experienced teachers in institutions that continually push for more summative assessments higher DOKs and rapid feedback while limiting preparation time As a teacher for over 25 years I know that any well outlined structured and comprehensive resource saves time in additional planning searching and preparing Use this book to help you identify and test students on topics that are important to their comprehension and success with their final test Chapter 1 Matter and change Chapter 2 measurement and calculations Chapter 3 Atoms The building blocks of matter Chapter 4 Arrangement of electrons in atoms Chapter 5 The periodic law Chapter 6 Chemical bonding Chapter 7 Chemical formulas and chemical compounds Chapter 8 Chemical equations and reactions Chapter 9 Stoichiometry Chapter 10 Physical characteristics of gases Chapter 11 Molecular composition of gases Chapter 12 Liquids and solids Chapter 13 Solutions Chapter 14 Ions in aqueous solution and colligative properties Chapter 15 Acids and bases Chapter 16 Acid base titrations Chapter 17 Reaction energy and reaction kinetics Chapter 18 Chemical equilibrium Chapter 19 Oxidation reduction reactions Chapter 20 Chemical thermodynamics Chapter 21 Carbon and hydrocarbons Chapter 22 Other organic compounds Chapter 23 Nuclear chemistry Chapter 24 Biological and Food chemistry Chapter 25 Environmental chemistry [A Stoichiometry Unit](#) David Callaghan, 2004 *Fast Track: Chemistry* The Princeton Review, 2021-03-23 GET UP TO SPEED WITH FAST TRACK CHEMISTRY Covering the most important material taught in high school chem class this essential review book breaks need to know content into accessible easily understood lessons Inside this book you ll find Clear concise summaries of the most important concepts terms and functions in chemistry Diagrams charts and graphs for quick visual reference Easy to follow content organization and illustrations With its friendly straightforward approach and a clean modern design crafted to appeal to visual learners this guidebook is perfect for catching up in class or getting ahead on exam review Topics covered in Fast Track Chemistry include Atomic structure Covalent bonding Intermolecular forces Stoichiometry Precipitation reactions Gas laws Thermochemistry Equilibrium and the solubility product constant Redox reactions Electrochemistry Acids and bases

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Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples

Solving Problems in Chemistry Rod O'Connor, Charles Mickey, Alton Hassell, 1977 **Chemistry 2007** David Wilson, 2006-12 Kaplan's guide includes 2 full length practice tests Diagnostic test to target areas for score improvement Detailed answer explanations Hundreds of practice questions from calculations of chemical equations to organic chemistry Explanations of important terms formulas and concepts Powerful strategies to help you score higher

The Ultimate General Chemistry Exam Guide Doctor Chemistry, 2017-09-03 This special addendum to the popular The Ultimate Guide to Learning or Teaching Chemistry This book contains the exam questions from a college AP chemistry class so you can see REAL test questions ahead of the exam The Ultimate Chemistry Series Teachers Never plan another lesson again Students Ace your upcoming exam This series covers all of the topics of High School Chemistry and General Chemistry including Accuracy and Significant Figures Mixtures Metric System Bonding Atomic Theory Periodic Table VSEPR Ionic and Covalent Bonding Geometric Bonding The Mole and

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