

ANSWER KEY

ATOMS AND ELECTRONS STUDY GUIDE

Learning Goals: (1) (10-11 P. 41) Explain a new scientific theory of atomic-scale phenomena (atomic theory) by describing changes in the atomic model over time, including those changes based on experimental observations. (10-11 P. 42) Explain the scientific theory of atomic-scale phenomena, including the scientific theories in terms of particles, motions, interactions, and the various forces that act to hold a system of these particles together, changes, and motions within the system. (10-11 P. 42) Develop properties of atoms and their motions in the scientific model for the composition of their systems, related to P. 42, as changes in the scientific knowledge of their first fundamental (scientific) properties, phenomena, and forces, such as forces, energy, mass, etc.

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Directions: Match the atomic theory selected on the correct card(s).

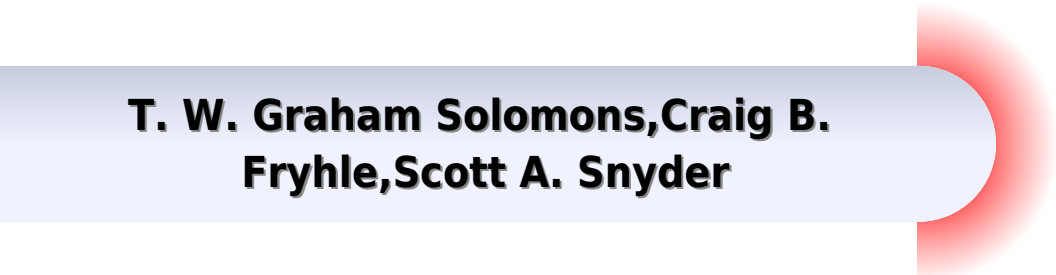
Model(s)	Model(s)	Atomic Theory (Atomic Model)
1C	Continuous (10-11 P. 41)	A. Everything is made of atoms (particles) that are small, divided, or separated.
1D	Discrete (10-11 P. 41)	B. Matter is made of particles (atoms) that are small, divided, or separated, and they are separated by spaces (voids) between them.
1E	Homogeneous (10-11 P. 41)	C. Matter is made of particles (atoms) that are small, divided, or separated.
1F	Heterogeneous (10-11 P. 41)	D. Matter is made of particles (atoms) that are small, divided, or separated.
1G	Mass (10-11 P. 41)	E. Matter is made of particles (atoms) that are small, divided, or separated.
1H	Energy (10-11 P. 41)	F. Matter is made of particles (atoms) that are small, divided, or separated.

Directions: Use the provided cards to answer the following questions:

- 1C** = 10-11 P. 41. The number of particles equals the number of electrons.
- 1D** = 10-11 P. 41. The number of particles is equal to the number of electrons.
- 1E** = 10-11 P. 41. The number of particles is equal to the number of electrons.

Energy And Electrons Study Guide Answers

**T. W. Graham Solomons, Craig B.
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