

# 2023 FINAL EXAM

## Grade 12 PAPER.1

# ECONOMICS



basic education

Department:  
Basic Education  
REPUBLIC OF SOUTH AFRICA

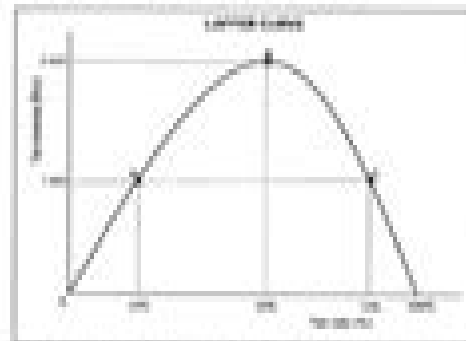
### SECTION 1: MULTIPLE CHOICE

### ANSWERS 1-10 (10 marks)

1.1 Answer the following questions.

- Q 1.1 Name the fifth saving motive of a business cycle. (2 x 1) 100  
Q 1.2 Why would a decrease in South African exports lead to the recession of the cycle? (2 x 1) 100

1.2 Study the graph below and answer the questions that follow.



- Q 1.1 Identify the first stage at which the government implements the strategy. 100  
Q 1.2 Name any form of interest rate used as a stimulant in the government. 100

2.2 Study the cartoon below and answer the questions that follow.



- 2.2.1 Identify a concept that relates to the changes in the consumer price index in the cartoon above. 100

# End Of Year Economics Paper1

**Tobias Bleicker**



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## **Table of Contents End Of Year Economics Paper1**

1. Understanding the eBook End Of Year Economics Paper1
  - The Rise of Digital Reading End Of Year Economics Paper1
  - Advantages of eBooks Over Traditional Books
2. Identifying End Of Year Economics Paper1
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an End Of Year Economics Paper1
  - User-Friendly Interface
4. Exploring eBook Recommendations from End Of Year Economics Paper1
  - Personalized Recommendations
  - End Of Year Economics Paper1 User Reviews and Ratings
  - End Of Year Economics Paper1 and Bestseller Lists

5. Accessing End Of Year Economics Paper1 Free and Paid eBooks
  - End Of Year Economics Paper1 Public Domain eBooks
  - End Of Year Economics Paper1 eBook Subscription Services
  - End Of Year Economics Paper1 Budget-Friendly Options
6. Navigating End Of Year Economics Paper1 eBook Formats
  - ePub, PDF, MOBI, and More
  - End Of Year Economics Paper1 Compatibility with Devices
  - End Of Year Economics Paper1 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of End Of Year Economics Paper1
  - Highlighting and Note-Taking End Of Year Economics Paper1
  - Interactive Elements End Of Year Economics Paper1
8. Staying Engaged with End Of Year Economics Paper1
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers End Of Year Economics Paper1
9. Balancing eBooks and Physical Books End Of Year Economics Paper1
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection End Of Year Economics Paper1
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine End Of Year Economics Paper1
  - Setting Reading Goals End Of Year Economics Paper1
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of End Of Year Economics Paper1
  - Fact-Checking eBook Content of End Of Year Economics Paper1
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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web  $m \times c \times \Delta t$  equation 1 where m is the total mass of the solution solute plus solvent c is the specific heat of the solution and  $\Delta t$  is the observed temperature change the specific heat of the solution is generally assumed to be the same as that of *guided inquiry designing the ideal hand warmer* - Dec 27 2022

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web oct 9 2021 sodium acetate coin or other small piece of metal plastic bags seal the coin inside a small bag place this bag inside a larger bag containing supersaturated sodium acetate get this by dissolving as much sodium acetate as possible in hot water this chemical hand warmer is reusable

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web the goal of my experiment and research is to evaluate the various number of compounds to be used as hand warmers and locate the ones that can raise the temperature to 20 degrees celsius in 100 g water and costs as little as possible

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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this experiment you will collect data that will allow you to calculate the heat of solution  $\Delta H_{\text{soln}}$  or the amount of energy released or absorbed in kJ/mol of the substance being dissolved

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web mar 4 2014 contains water the hand warmer is struck in a manner that ruptures the inner pouch releasing the ionic salt into the water of the outer pouch the salt dissolves and the water warms this type of hand warmer tends to produce a more vigorous heat than the dry powder type of hand warmer but does not produce heat for quite as long

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web hand warmers provide a unique and fun way to study the chemistry of crystal formation and heat generation by pressing a button in a pouch which contains a supercooled solution you start a rapid exothermic heat producing crystallization in this science fair project you will determine how the starting temperature affects hand warmer chemistry

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