

Chemistry Honors Worksheet: Formulas of a Compound

- moles The subscripts in a chemical formula represent the ratio of (moles / mass) in a compound
 - CO Which compound has the greatest percentage of carbon? (CO, CO₂, PbCO₃)
 - CH₄ What is the empirical formula of a compound that contains 0.2 moles of C and 0.8 moles of H
 - B₂g/mol Calculate the molecular mass of a compound if 0.02461 moles has a mass of 2.05 g.
 - Same The % composition of an empirical formula is (more / less / the same) than/as its molecular formula.
 - one How many moles of V are indicated in one mole of the compound V(C₂O₄)₂?
 - 1×10^{24} atoms A beaker of water contains 2×10^{24} atoms of H, how many atoms of O are also in the beaker?
 - 6.02×10^{23} 1 mole of a molecular compound has a mass of 165 grams. How many molecules are there?
 - 0% Fool's Gold is iron pyrite, FeS. What is the percentage of gold in a 1.0 gram sample of fool's gold?

- The pesticide DDT has been banned in the United States since 1962. This compound has a formula C₁₄H₉Cl₅. Answer all questions referring to DDT. Note: Use 35.5 g/mol for Cl 709.5 g/mol

- molecular Is this compound ionic or molecular?
- 5.2×10^{15} How many molecules of DDT are in a 6.1×10^{-4} g sample?
- 1.53×10^{17} atoms H How many atoms of H are in 2.00 mg of DDT?
- 7.513 Cl How many grams of Cl are contained in 10.0 grams of DDT?
- 7.84×10^{21} atoms C If a sample 5.04×10^{21} atoms of H, how many atoms of C would there also be?
- 2.97×10^{20} atoms C How many atoms of C are in a 25.0 mg of DDT?

- A compound of tantalum, Ta, and oxygen was produced in a lab by heating tantalum in a crucible. The data was collected.

Mass of crucible:	24.22 g		
Mass of crucible and tantalum:	26.14 g	<u>> 1.92g Ta</u>	<u>0.0106 mol Ta</u>
Mass of crucible and tantalum oxide:	26.56 g	<u>> 0.42g O</u>	<u>0.02625 mol O</u>

- Calculate the percent composition of the compound. 82.1% Ta 17.9% O
- What is the empirical formula of this compound? TaO_{2.5} - Ta₂O₅
- Is this compound ionic or covalent? ionic
- Name this compound: Tantalum(V) oxide * Need Roman Numerals!!!

- Determine the empirical formula of the following compounds:

a.) 63.0g Rb, 5.90g O



d.) 32.8% Cr, 67.2% Cl



b.) 4.95mg Th, 1.37mg S



e.) 58.0% Rb, 9.50% N, 32.5% O

c.) 2.13g Na, 2.32g As, 1.98g O



Empirical Molecular Formula Answers

JR Anderson



Empirical Molecular Formula Answers:

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