

Pre-lab Questions: The Determination of the Empirical Formula of a Hydrate

A hydrate is an ionic compound (ionic salt) that contains water molecules in its structure. An anhydrate (also called anhydrous salt) is the substance that remains after the water is removed from a hydrate. When a hydrate is heated, the water molecules are removed in the form of steam, leaving behind the anhydrate. To determine the formula of a hydrate experimentally, the mole ratio of the water portion to the mole ratio of the anhydrous portion must be calculated.

Questions:

During a Chem 121 experiment, a student obtained the following data after heating a sample of calcium chloride hydrate $\text{CaCl}_2 \cdot x\text{H}_2\text{O}$, where x refers to the number of water molecules hydrated per molecule of CaCl_2 .

Data:

Weight of crucible and lid (g)	32.1070
Weight of crucible, lid and hydrated salt (g)	36.8283
Weight of crucible, lid and anhydrous salt after	
a) 1 st heating (g)	35.7347
b) 2 nd heating (g)	35.6682
c) 3 rd heating (g)	35.6679

Be sure to show the general formula used for each step.

1. Use this information to determine the weight of the hydrated salt ($\text{CaCl}_2 \cdot x\text{H}_2\text{O}$):

a) the weight of the anhydrous salt (CaCl_2)

Note: the weight after the 2nd and 3rd heating are close enough in value to assume that constant weight has been reached and all the water molecules have been removed

b) the weight of water removed by heating

Empirical Formula Determination Lab Answers

M Carnoy



Empirical Formula Determination Lab Answers:

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