

Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.0 g of carbon and 6.67 g of hydrogen. Find the empirical formula.

$$40.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 3.33 \text{ mol C}$$

$$6.67 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.60 \text{ mol H}$$

C₃H₆

2. A compound is found to contain 20.0 g of carbon, 26.7 g of oxygen, and 6.67 g of hydrogen. Find the empirical formula.

$$20.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.67 \text{ mol C}$$

$$26.7 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 1.67 \text{ mol O}$$



$$6.67 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.60 \text{ mol H} \quad \frac{6.60 \text{ mol H}}{1.67 \text{ mol C}} = 4 \text{ H}$$

3. A compound is found to contain 40.0 g of carbon, 53.3 g of oxygen, and 6.67 g of hydrogen. Find the empirical formula.

$$40.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 3.33 \text{ mol C}$$



$$53.3 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 3.33 \text{ mol O}$$

$$6.67 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.60 \text{ mol H} \quad \frac{6.60 \text{ mol H}}{3.33 \text{ mol C}} = 2 \text{ H}$$

4. A 20.0 g compound is analyzed and found to contain 12.0 g carbon, 2.0 g hydrogen, and 6.0 g oxygen. Find the empirical formula.

$$12.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.00 \text{ mol C}$$



$$2.0 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 1.98 \text{ mol H}$$

$$6.0 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.375 \text{ mol O} \quad \frac{1.00 \text{ mol C}}{0.375 \text{ mol O}} = 2.67 \text{ C}$$

Empirical Formula Pogil Science Division Bellevue

Camilla Rothe



Empirical Formula Pogil Science Division Bellevue:

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