

Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.0 g of C and 6.67 g of H. Find the empirical formula.

$$40.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 3.33 \text{ mol C}$$

$$6.67 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.60 \text{ mol H}$$

C₃H₆

2. A compound is found to contain 20.0 g of Fe, 35.0 g of S, and 10.0 g of O. Find the empirical formula.

$$20.0 \text{ g Fe} \times \frac{1 \text{ mol Fe}}{55.85 \text{ g Fe}} = 0.358 \text{ mol Fe}$$

$$35.0 \text{ g S} \times \frac{1 \text{ mol S}}{32.07 \text{ g S}} = 1.09 \text{ mol S}$$

$$10.0 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.625 \text{ mol O}$$



3. A compound is found to contain 40.0 g of C, 6.67 g of H, and 13.33 g of O. Find the empirical formula.

$$40.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 3.33 \text{ mol C}$$

$$6.67 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.60 \text{ mol H}$$

$$13.33 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.833 \text{ mol O}$$



4. A compound is found to contain 40.0 g of C, 6.67 g of H, and 13.33 g of O. Find the empirical formula.

$$40.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 3.33 \text{ mol C}$$

$$6.67 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.60 \text{ mol H}$$

$$13.33 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.833 \text{ mol O}$$



Empirical Formula Pogil Answers

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