

# Empirical Formula and Molecular Formula

Lab 10: Molecular Formulas and Empirical Formulas

## Lab 10

Students will learn how to determine the empirical formula of a compound from its mass percent composition. They will also learn how to determine the molecular formula of a compound from its mass percent composition and its molar mass. This activity will require the use of a calculator.

### Model 1: Comparison of Percent Composition and Empirical Formula

Compound	Empirical Formula	Molecular Formula	Empirical Formula	Percent Composition (Carbon)	Percent Composition (Hydrogen)
1	ethane	$C_2H_6$	$CH_3$		
2	propane	$C_3H_8$	$CH_2.67$		
3	ethylene (acetylene)	$C_2H_2$	$CH$		
4	benzene	$C_6H_6$	$CH$		

1. Looking at the table, what do you notice about the percent composition of the compounds? (Refer to the Model 1. What do you notice about the percent composition of the compounds?)
2. Identify the compounds in Model 1 that have the same percent composition.
3. Using the compounds in Model 1, suggest an explanation for why compounds with different molecular formulas but the same percent composition have the same percent composition.
4. For each substance in Model 1, you will notice that the molecular formula and empirical formula are multiples of each other. (For example, ethane has the molecular formula  $C_2H_6$  and the empirical formula  $CH_3$ .)

# Empirical Formula And Molecular Formula Pogil Answers

**Victor M. Corman**



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