

Empirical and Molecular Formula Practice KEY

Empirical Formula

1. 36.5% Na, 25.4% S, and 38.1% O forms this compound.

$$\text{Na} - 36.5/23 = 1.56/0.8 = 2$$

$$\text{S} - 25.4/32 = 0.8/0.8 = 1$$

$$\text{O} - 38.1/16 = 2.4/0.8 = 3$$



2. A compound made of 72.2% Magnesium and 27.8% nitrogen by mass.

$$\text{Mg} - 72.2/24.3 = 3/2 = 1.5(2) = 3$$

$$\text{N} - 27.8/14 = 2/2 = 1(2) = 2$$



3. A chromium compound with 68.4% Cr and 31.6% O.

$$\text{Cr} - 68.4/52 = 1.315/1.315 = 1(2) = 2$$

$$\text{O} - 31.6/16 = 1.975/1.315 = 1.5(2) = 3$$



4. A compound with the composition of 24% C, 7% H, 38% F, and 31% P.

$$\text{C} - 24/12 = 2/1 = 2$$

$$\text{H} - 7/1 = 7/1 = 7$$

$$\text{F} - 38/19 = 2/1 = 2$$

$$\text{P} - 31/31 = 1/1 = 1$$



5. 63.5% silver, 8.2% nitrogen, and 28.3% oxygen make up this compound.

$$\text{Ag} - 63.5/107.87 = 0.589/0.585 = 1$$

$$\text{N} - 8.2/14.01 = 0.585/0.585 = 1$$

$$\text{O} - 28.3/16 = 1.769/0.585 = 3$$



6. 8.81g C and 91.2g Chlorine forms this compound.

$$\text{C} - 8.81/12 = 0.73/0.73 = 1(2) = 2$$

$$\text{Cl} - 91.2/35 = 2.61/0.73 = 3.5(2) = 7$$



7. A compound in a soft drink that has 0.31g hydrogen, 3.2g phosphorus, and 6.5g oxygen.

$$\text{H} - 0.31/1 = 0.31/.1 = 3.1$$

$$\text{P} - 3.2/31 = 0.10/.1 = 1$$

$$\text{O} - 6.5/16 = 0.40/.1 = 4$$



Empirical And Molecular Formula Key

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Table of Contents Empirical And Molecular Formula Key

1. Understanding the eBook Empirical And Molecular Formula Key
 - The Rise of Digital Reading Empirical And Molecular Formula Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Empirical And Molecular Formula Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Empirical And Molecular Formula Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical And Molecular Formula Key
 - Personalized Recommendations
 - Empirical And Molecular Formula Key User Reviews and Ratings
 - Empirical And Molecular Formula Key and Bestseller Lists
5. Accessing Empirical And Molecular Formula Key Free and Paid eBooks
 - Empirical And Molecular Formula Key Public Domain eBooks
 - Empirical And Molecular Formula Key eBook Subscription Services
 - Empirical And Molecular Formula Key Budget-Friendly Options
6. Navigating Empirical And Molecular Formula Key eBook Formats

- ePub, PDF, MOBI, and More
- Empirical And Molecular Formula Key Compatibility with Devices
- Empirical And Molecular Formula Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Empirical And Molecular Formula Key
 - Highlighting and Note-Taking Empirical And Molecular Formula Key
 - Interactive Elements Empirical And Molecular Formula Key
- 8. Staying Engaged with Empirical And Molecular Formula Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Empirical And Molecular Formula Key
- 9. Balancing eBooks and Physical Books Empirical And Molecular Formula Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Empirical And Molecular Formula Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Empirical And Molecular Formula Key
 - Setting Reading Goals Empirical And Molecular Formula Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Empirical And Molecular Formula Key
 - Fact-Checking eBook Content of Empirical And Molecular Formula Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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