

EXPERIMENT: INVESTIGATING THE DENSITY OF WATER AND ICE

Syllabus reference 8.4.1

INTRODUCTION

There are a number of important properties that assist in identifying elements and pure compounds. These include density, melting point and boiling point.

The density of a substance is the measure of its compactness and is calculated from the ratio:

$$\text{density} = \frac{\text{mass (g)}}{\text{volume (mL)}}$$

In this experiment you will:

- a measure the mass of several different volumes of water and use these to calculate the density of water
- b use a similar procedure to calculate the density of ice.

EQUIPMENT

- 100 mL beaker
- wide diameter measuring cylinder
- electronic balance
- distilled water
- ice cubes (made with distilled water and small enough to fit into the measuring cylinder)
- 10 mL pipette
- forceps or glass rod
- thermometer

Part A: Density of water

PROCEDURE

- 1 Accurately measure the mass of the empty beaker. Record this in the table below.
- 2 Add a known volume of distilled water (use the pipette) to the beaker and reweigh. Record this in the table.
- 3 Measure the temperature of the water.
- 4 Repeat this procedure until you have at least 5 measurements (use 10 mL increments from the starting volume). Record all measurements in the table.

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Table of Contents Mcgraw Hill Conquering Chemistry Prelim Answers

1. Understanding the eBook Mcgraw Hill Conquering Chemistry Prelim Answers
 - The Rise of Digital Reading Mcgraw Hill Conquering Chemistry Prelim Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Conquering Chemistry Prelim Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Conquering Chemistry Prelim Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Conquering Chemistry Prelim Answers
 - Personalized Recommendations
 - Mcgraw Hill Conquering Chemistry Prelim Answers User Reviews and Ratings
 - Mcgraw Hill Conquering Chemistry Prelim Answers and Bestseller Lists

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 - Mcgraw Hill Conquering Chemistry Prelim Answers Public Domain eBooks
 - Mcgraw Hill Conquering Chemistry Prelim Answers eBook Subscription Services
 - Mcgraw Hill Conquering Chemistry Prelim Answers Budget-Friendly Options
6. Navigating Mcgraw Hill Conquering Chemistry Prelim Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Mcgraw Hill Conquering Chemistry Prelim Answers Compatibility with Devices
 - Mcgraw Hill Conquering Chemistry Prelim Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcgraw Hill Conquering Chemistry Prelim Answers
 - Highlighting and Note-Taking Mcgraw Hill Conquering Chemistry Prelim Answers
 - Interactive Elements Mcgraw Hill Conquering Chemistry Prelim Answers
8. Staying Engaged with Mcgraw Hill Conquering Chemistry Prelim Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Conquering Chemistry Prelim Answers
9. Balancing eBooks and Physical Books Mcgraw Hill Conquering Chemistry Prelim Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Conquering Chemistry Prelim Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mcgraw Hill Conquering Chemistry Prelim Answers
 - Setting Reading Goals Mcgraw Hill Conquering Chemistry Prelim Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mcgraw Hill Conquering Chemistry Prelim Answers
 - Fact-Checking eBook Content of Mcgraw Hill Conquering Chemistry Prelim Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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