

- 1.16.** 1 lb mol = 453.59 g mol (App. 1). At 1 atm, 20°C,  
Number of molecules of oxygen, from Avogadro's number:

$$0.209 \times 6.022 \times 10^{23} \times 453.59 = 5.71 \times 10^{25}$$

Number is not affected by temperature or pressure. At  
1 atm and 30°C, in 1 kg mol the number is:

$$0.209 \times 6.022 \times 10^{23} \times 1000 = 1.259 \times 10^{26}$$

- 1.17.** The internal energy  $E$  of an object is the product of its  
mass and the sum of its kinetic energy, potential energy,  
and enthalpy. The *change* in internal energy,  $\Delta E$ , therefore  
equals the left-hand side of Eq. (1.59), and we can say,

$$\Delta E = Q - W$$

# Mccabe Solutions Manual

**CO Houle**



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