

Chapter 3: Matter and Energy – Study Guide

*Do not forget about the extra resources available to help you study and prepare for this test!

- Textbook pages 81- 93 (answers to odd questions in the back of the book)
- Mastering Chemistry Study Area: <http://www.pearsonmylabandmastering.com/northamerica/>
- Miss Marnik's Website: <http://www.northallegheny.org/Page/16665>
- Labs, problem packets, and homework assignments

Section 1 – Classifying Matter: (text pages 57-62)

Need to be able to:

- Define matter, atoms, molecules, element, compound, and mixture.
- Classify matter as and distinguish between solid, liquid, or gas.
- Classify matter as element, compound, or mixture.

Practice Problems:

1. Match the following terms with the phrases below:

- | | |
|--|--------------------------|
| a. State of matter in which atoms or molecules pack close to each other in fixed locations | 1. heterogeneous mixture |
| b. State of matter in which atoms or molecules are packed close together but are free to move around and by each other | 2. homogeneous mixture |
| c. Substance composed of only one type of atom or molecule | 3. liquid |
| d. Matter composed of two or more different types of atoms or molecules combined in variable proportions | 4. mixture |
| e. Mixture that has two or more regions with different compositions | 5. pure substance |
| f. Mixture that has the same composition throughout | 6. solid |

2. Which of the following can be classified as matter?

- | | | | |
|----------|------------|---------------------|-----------|
| a. atoms | c. wind | e. potential energy | g. carbon |
| b. gas | d. liquids | f. temperature | h. light |

3. Classify each type of matter as a pure substance or a mixture. If it is a pure substance, classify it as an element or a compound. If it is a mixture, classify it as homogeneous or heterogeneous.

- | | |
|----------------------------------|---------------|
| a. limestone (CaCO_3) | e. Chex mix |
| b. orange juice with pulp | f. Table salt |
| c. helium | g. Brass |
| d. Pacific ocean | h. oxygen |

Matter And Energy Study Guide

AN Whitehead



Matter And Energy Study Guide:

Matter and Energy ,2005 **Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J.

Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. *Matter and Energy* ,2005 *Study Guide to Accompany Basics for Chemistry* Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18-chapter text designed to be used with Basics for Chemistry textbook. Each chapter contains Overview, Topical Outline, Skills, and Common Mistakes, which are all keyed to the textbook for easy cross-reference. The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms, a summary of general concepts, and a list of numerical exercises, while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook. The Fill-in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter, and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples. The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry, while the Practical Test section includes matching and multiple-choice questions that comprehensively cover almost every concept and numerical problem in the chapter. After briefly dealing with an overview of chemistry, this book goes on exploring the concept of matter, energy, measurement, problem solving, atom, periodic table, and chemical bonding. These topics are followed by discussions on writing names and formulas of compounds, chemical formulas, and the mole, chemical reactions, calculations based on equations, gases, and the properties of a liquid. The remaining chapters examine the solutions, acids, bases, salts, oxidation-reduction reactions, electrochemistry, chemical kinetics, and equilibrium and nuclear, organic, and biological chemistry. This study guide will be of great value to chemistry teachers and students. Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon: buckyballs, graphite,

and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement.

Chemical Principles Study Guide/Solutions Manual John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18. Written for general chemistry courses. Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications. *Chemical Principles Student's Study Guide & Solutions Manual* John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19. This combination manual is designed to help students avoid common mistakes and understand the material better. The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text. *Study Guide/Selected Solutions Manual* Julie R. Frentrop, 2002-06. Contains a brief overview of every chapter, review of skills, self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book.

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