

Quantitative analysis

Read through the whole method before starting any practical work. Where appropriate, prepare a table for your results in the space provided.

Show your working and appropriate significant figures in the final answer to **each** step of your calculations.

- 1 In this experiment you will carry out a titration to determine the relative formula mass, M_r , of a monoprotic acid, HX.

FA 1 is HX, the monoprotic acid.

FA 2 is $0.0450 \text{ mol dm}^{-3}$ sodium carbonate, Na_2CO_3 .
methyl orange indicator

(a) Method

Preparing a solution of FA 1

- Weigh the empty 250 cm^3 beaker. Record the mass.
- Transfer all the **FA 1** into the beaker.
- Weigh the beaker and **FA 1**. Record the mass.
- Calculate and record the mass of **FA 1** used.
- Add approximately 100 cm^3 of distilled water to the **FA 1** in the beaker.
- Stir the mixture with a glass rod until all the **FA 1** has dissolved.
- Transfer this solution into the 250 cm^3 volumetric flask.
- Wash the beaker with distilled water and transfer the washings to the volumetric flask.
- Rinse the glass rod with distilled water and transfer the washings to the volumetric flask.
- Make up the solution in the volumetric flask to the mark using distilled water.
- Shake the flask thoroughly.
- This solution of HX is **FA 3**. Label the flask **FA 3**.

Results

Titration

- Fill the burette with **FA 3**.
- Pipette 25.0 cm^3 of **FA 2** into a conical flask.
- Add several drops of methyl orange indicator.
- Perform a rough titration and record your burette readings in the space below.

The rough titre is cm^3 .

June 2013 Chemistry 9701 Paper 35 Question

Minjie Lin



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Table of Contents June 2013 Chemistry 9701 Paper 35 Question

1. Understanding the eBook June 2013 Chemistry 9701 Paper 35 Question
 - The Rise of Digital Reading June 2013 Chemistry 9701 Paper 35 Question
 - Advantages of eBooks Over Traditional Books
2. Identifying June 2013 Chemistry 9701 Paper 35 Question
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an June 2013 Chemistry 9701 Paper 35 Question
 - User-Friendly Interface
4. Exploring eBook Recommendations from June 2013 Chemistry 9701 Paper 35 Question
 - Personalized Recommendations
 - June 2013 Chemistry 9701 Paper 35 Question User Reviews and Ratings
 - June 2013 Chemistry 9701 Paper 35 Question and Bestseller Lists
5. Accessing June 2013 Chemistry 9701 Paper 35 Question Free and Paid eBooks
 - June 2013 Chemistry 9701 Paper 35 Question Public Domain eBooks

- June 2013 Chemistry 9701 Paper 35 Question eBook Subscription Services
- June 2013 Chemistry 9701 Paper 35 Question Budget-Friendly Options
- 6. Navigating June 2013 Chemistry 9701 Paper 35 Question eBook Formats
 - ePub, PDF, MOBI, and More
 - June 2013 Chemistry 9701 Paper 35 Question Compatibility with Devices
 - June 2013 Chemistry 9701 Paper 35 Question Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of June 2013 Chemistry 9701 Paper 35 Question
 - Highlighting and Note-Taking June 2013 Chemistry 9701 Paper 35 Question
 - Interactive Elements June 2013 Chemistry 9701 Paper 35 Question
- 8. Staying Engaged with June 2013 Chemistry 9701 Paper 35 Question
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers June 2013 Chemistry 9701 Paper 35 Question
- 9. Balancing eBooks and Physical Books June 2013 Chemistry 9701 Paper 35 Question
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection June 2013 Chemistry 9701 Paper 35 Question
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine June 2013 Chemistry 9701 Paper 35 Question
 - Setting Reading Goals June 2013 Chemistry 9701 Paper 35 Question
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of June 2013 Chemistry 9701 Paper 35 Question
 - Fact-Checking eBook Content of June 2013 Chemistry 9701 Paper 35 Question
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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