

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Identify a conjugate acid-base pair in chemical reactions involving acids and bases.
2. Calculate the pH of a solution from the concentration of H^+ or H_3O^+ ions, and vice versa.
3. Determine the pH of a solution from K_a and K_b .

Vocabulary

acid	1. conjugate base	acid solution
acid-base indicator	2. hydronium ion	acidic solution
acidic solution	3. autoionization	pH

Key Equation

- a. $pH = -\log [H^+]$ (ppm 15.2 and 15.3)
- b. $[H^+] = 10^{-pH}$ (ppm 15.2 and 15.3)

Part A Completion

Write the unknown number in each blank space. (ppm 15.2 and 15.3)

1. The K_a of a weak acid is 1.5×10^{-4} . The K_b of its conjugate base is _____.
2. The K_b of a weak base is 1.5×10^{-4} . The K_a of its conjugate acid is _____.
3. The pH of a solution is 4.5. The concentration of H^+ ions is _____ M.
4. The pH of a solution is 10.5. The concentration of OH^- ions is _____ M.
5. The pH of a solution is 4.5. The concentration of OH^- ions is _____ M.
6. The pH of a solution is 10.5. The concentration of H^+ ions is _____ M.
7. The pH of a solution is 4.5. The concentration of H^+ ions is _____ M.
8. The pH of a solution is 10.5. The concentration of OH^- ions is _____ M.
9. The pH of a solution is 4.5. The concentration of OH^- ions is _____ M.
10. The pH of a solution is 10.5. The concentration of H^+ ions is _____ M.

Hydrogen Ions And Acidity Section Review Answers

David Baud



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