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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al $\times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol Al}} \times 101.96 \text{ g/mol} = 63.2 \text{ g}$ c determine the percent yield $\frac{26.98 \text{ g}}{63.2 \text{ g}} \times 100 = 42.7\%$

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