

GCSE Chemistry Trilogy Glossary.



4.1 Atomic structure and the periodic table

4.1.1 A simple model of the atom, symbols, relative atomic mass, electronic charge and isotopes

Key Term	Definition	Example
Atom	The smallest building block of an element, with a central nucleus, made up of protons and neutrons. There are orbiting electrons around the nucleus.	The copper spoon is made up of a very large number of copper atoms.
Symbol	Letters that relate to the name of the element, but the names used can be based on old Greek or Latin names.	Na is the symbol of the metal Sodium, the symbol is short for Νάτριο, its Greek name.
Charge	A characteristic of many subatomic particles such as the electron and proton that affects their interaction with an electric or magnetic field.	Protons are a subatomic particle that have a positive charge.
Periodic	An effect or characteristics that repeats after a given time or given conditions.	The rows in the periodic table start once atoms in the period above have filled their outer shell of electrons.
Element	A substance with the same type of atoms.	Pure gold is an element as it only contains gold atoms.
Compound	A substance with two or more atoms, these can be from different elements, and these are chemically bonded.	Two or more atoms that need a lot of energy to be separated, make up a compound.
Mixture	A substance made of two or more different elements or compounds that can be easily separated.	The salad dressing separated into its oil and water component easily.
Chromatography	A separation technique that pull apart the substances in a mixture due to the different absorption properties between paper and the substance.	The inks used to make water soluble black ink were separated into 3 unexpected colours, using filter paper and water.
Distillation	A separation technique that keeps apart substances from a mixture due to their different boiling points.	The rose oil evaporated before the water contained in the petals, in the distillation process for making perfume.
Crystallisation	The formation of regularly shaped structures from a salt solution. It is a separation technique to obtain the solid salt from its solution.	There was a white and fine solid left in the container of Sea water after a few days, and once all the water had evaporated.

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web a similar unit of concentration is molality m which is defined as the number of moles of solute per kilogram of solvent not per liter of solution

$$\text{molality} = \frac{\text{moles solute}}{\text{kilograms solvent}}$$
 mathematical manipulation of molality is the

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web 1 molarity M mass molar mass $\times$ 0 4500 l 0 825 g 141 9579 g mol $\times$ 0 0129 m 2 molality 0 825 g 141 9579 g mol 0 00581158 mol 0 00581158 mol 0 4500 kg 0 0129 m 3 mole fraction n_2 hpo 4 0 825 g 141 9579 g mol 0 00581158 mol h 2 o 450 0 g 18 015 g mol 24 97918401 mol

16 11 molality chemistry libretexts - Mar 16 2023

web molality differs from molarity only in the denominator while molarity is based on the liters of solution molality is based on the kilograms of solvent concentrations expressed in molality are used when studying properties of solutions related to vapor pressure and temperature changes

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web what mass of CoCl_2 in grams is needed for the solution express the answer using 3 significant figures

8 3 concentrations of solutions problems chemistry libretexts - Dec 13 2022

web the hardness of water hardness count is usually expressed in parts per million by mass of CaCO_3 which is equivalent to milligrams of CaCO_3 per liter of water what is the molar concentration of Ca^{2+} ions in a water sample with a hardness count of 175 mg CaCO_3 l

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web answer the mass percentage of an element of a solution is defined as the mass of the solute in grammes present in 100 g of the solution it s written like this

$$\text{begin array l} \frac{\text{mass of a component}}{\text{mass of component in solution total mass of solution}} \times 100 \text{ end array}$$

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web aug 13 2017 molarity practice problems this chemistry video tutorial explains how to solve common molarity problems it discusses how to calculate the concentration of a solution given the mass in grams