

QUESTION 3

3. 1.Learners made a mini volcano in science laboratory by adding sodium bicarbonate to ethanoic acid. They added 100 ml of a $0,2 \text{ mol}\cdot\text{dm}^{-3}$ ethanoic acid solution to 10 g of NaHCO_3 to start the reaction of the volcano.

The balanced equation for this reaction is:



- 3.1.1 Define the term limiting reagent. (2)
- 3.1.2 Determine the limiting reagent in this reaction. (6)
- 3.1.3 Calculate the mass of the other substance in excess. (3)
- 3.1.4 Calculate the volume of CO_2 produced at STP. (4)

3.2 A solution of potassium hydroxide (KOH) is prepared by dissolving 3,36 g crystals of KOH in 250 cm^3 of water.

Calculate the concentration of the potassium hydroxide solution. (4)

3.3 25 cm^3 of a potassium hydroxide solution of concentration $0,25 \text{ mol}\cdot\text{dm}^{-3}$ completely neutralises a dilute solution of sulphuric acid (H_2SO_4) in a flask.

The unbalanced equation below represents the reaction that takes place:



- 3.3.1 Balance the above equation. (2)
- 3.3.2 Calculate the mass of sulphuric acid in the flask. (5)

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Table of Contents Grade11 Term 3 Physical Science Test Consists Of What

1. Understanding the eBook Grade11 Term 3 Physical Science Test Consists Of What
 - The Rise of Digital Reading Grade11 Term 3 Physical Science Test Consists Of What
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade11 Term 3 Physical Science Test Consists Of What
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade11 Term 3 Physical Science Test Consists Of What
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade11 Term 3 Physical Science Test Consists Of What
 - Personalized Recommendations
 - Grade11 Term 3 Physical Science Test Consists Of What User Reviews and Ratings
 - Grade11 Term 3 Physical Science Test Consists Of What and Bestseller Lists
5. Accessing Grade11 Term 3 Physical Science Test Consists Of What Free and Paid eBooks

- Grade11 Term 3 Physical Science Test Consists Of What Public Domain eBooks
 - Grade11 Term 3 Physical Science Test Consists Of What eBook Subscription Services
 - Grade11 Term 3 Physical Science Test Consists Of What Budget-Friendly Options
6. Navigating Grade11 Term 3 Physical Science Test Consists Of What eBook Formats
- ePub, PDF, MOBI, and More
 - Grade11 Term 3 Physical Science Test Consists Of What Compatibility with Devices
 - Grade11 Term 3 Physical Science Test Consists Of What Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Grade11 Term 3 Physical Science Test Consists Of What
 - Highlighting and Note-Taking Grade11 Term 3 Physical Science Test Consists Of What
 - Interactive Elements Grade11 Term 3 Physical Science Test Consists Of What
8. Staying Engaged with Grade11 Term 3 Physical Science Test Consists Of What
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade11 Term 3 Physical Science Test Consists Of What
9. Balancing eBooks and Physical Books Grade11 Term 3 Physical Science Test Consists Of What
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade11 Term 3 Physical Science Test Consists Of What
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Grade11 Term 3 Physical Science Test Consists Of What
- Setting Reading Goals Grade11 Term 3 Physical Science Test Consists Of What
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Grade11 Term 3 Physical Science Test Consists Of What
- Fact-Checking eBook Content of Grade11 Term 3 Physical Science Test Consists Of What
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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