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Anne Eadie and Gretel Lampe

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# Grade 11 Math Study Guide

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calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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g  $\text{Al}$   $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$  4 mol  $\text{Al}$  1 mol  $\text{Al}_2\text{O}_3$  c determine the percent yield  
yield actual  $\times 100$  260.0 g  $\times 100$  86.01 theoretical 302.3 g

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c is the limiting reagent 4.33 g of h 2 are left over 13

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 $\text{CH}_3\text{OH}$  can be produced by the following reaction  $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$  calculate the theoretical yield of  $\text{CH}_3\text{OH}$  if 68.5 g of  $\text{CO}$  is  
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