

## UNIT III - THERMOCHEMISTRY

1. Calculate the H-F bond energy given the following data.



$$\Delta H_{\text{rxn}} = -544 \text{ kJ}$$

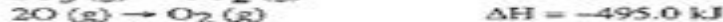
- A. 1130 kJ/mol    B. 565 kJ/mol    **C. 283 J/mol**    D. 300 kJ/mol    E. 134 kJ/mol
2. 1.0 g of the rocket fuel hydrazine ( $\text{N}_2\text{H}_4$ ) is burned in a bomb calorimeter. The temperature of the bomb calorimeter changes from 24.63 °C to 28.16 °C. The heat capacity of the bomb calorimeter is 837 J/K. Calculate the heat of combustion per mol of hydrazine in the bomb calorimeter.

- A. 95.0 kJ/mol**    B. 2.96 kJ/mol    C.  $3.55 \times 10^3$   
D. 92.4 J/mol    E.  $1.13 \times 10^5$  kJ/mol

3. Calculate the enthalpy change for the following reaction of nitric oxide NO with ozone
- $\text{O}_3$
- .



Given the following data:



- A. -1083.7 kJ    B. -693.8 kJ    C. -198.8 kJ    **D. -93.7 kJ**    E. -93.7 kJ
4. An endothermic reaction does no work on the surroundings. What values do you predict for the heat and energy change of this reaction?

- A.  $q < 0$ ,  $\Delta E > 0$     **B.  $q < 0$ ,  $\Delta E < 0$**     C.  $q > 0$ ,  $\Delta E > 0$   
D.  $q > 0$ ,  $\Delta E < 0$     E.  $q > 0$ ,  $\Delta E = 0$

5. M&M candies contain 209.0 Calories per 1.5 oz serving. The fat content accounts for 38.3% of these Calories. What is the energy of the fat content in kJ?

- A. 19.1 kJ    B. 874 kJ    **C.  $1.91 \times 10^{-2}$  kJ**  
D.  $3.35 \times 10^5$  kJ    E. 335 kJ

6. Calculate the heat released when 2.00 g of butane from a cigarette lighter is burned.



- A. -90.7 kJ    **B. -98.3 kJ**    C. -49.2 kJ    D. -108.4 kJ    E. -196.6 kJ
7. When Karl Kaverman adds chilled grog to a granite mug, 10.9 kJ of energy is transferred from the mug to the grog. If the mug has a mass of 625 g and an initial temperature of 25 °C, what is the final temperature of the mug? The specific heat of granite is 0.79 J/g·K.

- A. 2.9 °C    B. 14 °C    **C. 22 °C**    D. 25 °C    E. 47 °C

# General Chemistry Exam Questions And Answers

**Kaplan**



## **General Chemistry Exam Questions And Answers:**

**General Chemistry Workbook** Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

**100 Questions (and Answers) About Tests and Measurement** Bruce B. Frey, 2014-04-03 100 Questions and Answers About Tests and Measurement asks and answers important questions about the world of social science measurement It is ideal as an introduction to students new to the concepts to advanced students and professionals looking to review ideas and procedures as well as to those interested in knowing more about a test they have to take or how to interpret the score they receive

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**Cracking College Admissions** Princeton Review,2004 Presents a comprehensive guide to assisting with the college admissions process and includes information on choosing the right school creating schedules understanding tuition and financial aid and making a good first impression on the admissions committee

**Excel Preliminary Information Processes and Technology** Mary O'Connor-Nickel,2002

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