

## Thermochemistry and Gas Laws Unit Review

### Energy Math

1. How much energy is released when a brick with a mass of 329g is cooled from 80°C to 50°C? The specific heat capacity of brick is 0.84 J/g°C.

$$Q = mc\Delta T$$

$$(329g)(0.84J/g^{\circ}C)(-30^{\circ}C)$$

$$Q = -8290J$$

2. What is the mass of a hotdog that releases 3600 J of energy as it is cooled from 95°C to 45 °C? The specific heat capacity of a hotdog is 1.5 J/g°C.

$$m = \frac{-3600J}{(1.5J/g^{\circ}C)(-50^{\circ}C)} = 48g$$

3. What is the specific heat capacity of a 350 g object that rises in temperature 15°C when 55,900 J of energy are added?

$$C = \frac{55900J}{(350g)(15^{\circ}C)} = 10.6J/g^{\circ}C$$

4. 45,920 J of energy are added to a 365g block of copper originally at 35°C. What will the new temperature of the copper be? The specific heat capacity of copper is 0.385 J/g°C.

$$\Delta t = \frac{45920J}{(365g)(0.385J/g^{\circ}C)} = 326^{\circ}C = T_f - 35^{\circ}C$$

$$+35^{\circ}C \quad +35^{\circ}C$$

$$361^{\circ}C = T_f$$

5. 89,290 J of energy are removed from a 670 g block of aluminum originally at 125°C. What will the new temperature of the aluminum be? The specific heat capacity of aluminum is 0.902 J/g°C.

$$\Delta t = \frac{-89290J}{(670g)(0.902J/g^{\circ}C)} = -148^{\circ}C = T_f - 125^{\circ}C$$

$$+125^{\circ}C \quad +125^{\circ}C$$

$$-23^{\circ}C = T_f$$

# Energy And Thermochemistry Review Guide Key

**Patrick Vollmar**



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## **Table of Contents Energy And Thermochemistry Review Guide Key**

1. Understanding the eBook Energy And Thermochemistry Review Guide Key
  - The Rise of Digital Reading Energy And Thermochemistry Review Guide Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Energy And Thermochemistry Review Guide Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Energy And Thermochemistry Review Guide Key
  - User-Friendly Interface

4. Exploring eBook Recommendations from Energy And Thermochemistry Review Guide Key
  - Personalized Recommendations
  - Energy And Thermochemistry Review Guide Key User Reviews and Ratings
  - Energy And Thermochemistry Review Guide Key and Bestseller Lists
5. Accessing Energy And Thermochemistry Review Guide Key Free and Paid eBooks
  - Energy And Thermochemistry Review Guide Key Public Domain eBooks
  - Energy And Thermochemistry Review Guide Key eBook Subscription Services
  - Energy And Thermochemistry Review Guide Key Budget-Friendly Options
6. Navigating Energy And Thermochemistry Review Guide Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Energy And Thermochemistry Review Guide Key Compatibility with Devices
  - Energy And Thermochemistry Review Guide Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Energy And Thermochemistry Review Guide Key
  - Highlighting and Note-Taking Energy And Thermochemistry Review Guide Key
  - Interactive Elements Energy And Thermochemistry Review Guide Key
8. Staying Engaged with Energy And Thermochemistry Review Guide Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Energy And Thermochemistry Review Guide Key
9. Balancing eBooks and Physical Books Energy And Thermochemistry Review Guide Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Energy And Thermochemistry Review Guide Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Energy And Thermochemistry Review Guide Key
  - Setting Reading Goals Energy And Thermochemistry Review Guide Key
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Energy And Thermochemistry Review Guide Key
  - Fact-Checking eBook Content of Energy And Thermochemistry Review Guide Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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