

Name Key Period \_\_\_\_\_

**Worksheet: Calculating Empirical & Molecular Formulas**

1. The empirical formula for the compound having the formula  $H_4C_2O_4$  is:

- [A]  $C_2H_2$  [B]  $CO_2H$  [C]  $COH$  [D]  $C_2O_4H_4$  [E]  $COH_2$

2. Calculate the empirical formula of a compound that is 85.6% C and 14.4% H (by mass).

- [A]  $CH_2$  [B]  $CH$  [C]  $C_2H_4$  [D]  $C_2H_6$  [E]  $C_3H_8$

$$C: \frac{85.6}{12.01} = 7.12 \quad H: \frac{14.4}{1.01} = 14.25 = \frac{14.25}{7.12} = 2$$

3. A compound is analyzed and found to contain 12.1% carbon, 56.2% oxygen, and 71.7% chlorine (by mass). Calculate the empirical formula of this compound.

- [A]  $COCl_2$  [B]  $CO_2Cl_2$  [C]  $CO_2Cl$  [D]  $COCH$  [E]  $COCl$

$$C: \frac{12.1}{12.01} = 1 \quad O: \frac{56.2}{16.00} = 3.51 \quad Cl: \frac{71.7}{35.45} = 2$$

4. A compound contains 40.0% carbon, 6.7% hydrogen, and 53.3% oxygen (by mass). Calculate the empirical formula.

- [A]  $C_2H_2O$  [B]  $CH_2O$  [C]  $CH_4O$  [D]  $C_2H_2O_2$  [E]  $C_2H_4O_2$

$$C: \frac{40.0}{12.01} = 3.33 \quad H: \frac{6.7}{1.01} = 6.63 \quad O: \frac{53.3}{16.00} = 3.33 = \frac{3.33}{3.33} = 1$$

5. A compound contains 25.94% N and 74.06% O (by mass). What is the empirical formula?

$$\frac{25.94g N}{14.01} = \frac{1.85}{1.84} = 1 \times 2 = 2$$

$$\frac{74.06g O}{16.00} = \frac{4.63}{1.84} = 2.5 \times 2 = 5$$



6. Determine the empirical formula of a compound containing 54.2% F and 45.8% S (by mass).

$$\frac{54.2g F}{19.00g/mol} = \frac{2.85}{1.43} mol F = 1.99 \sim 2$$

$$\frac{45.8g S}{32.07g/mol} = \frac{1.43}{1.43} mol S = 1$$



7. A compound has 40.68% carbon, 5.12% hydrogen, and 54.20% oxygen (by mass). Calculate its empirical formula.

$$\frac{40.68g C}{12.01g/mol} = \frac{3.387}{3.387} \quad \frac{5.12g H}{1.01g/mol} = \frac{5.069}{3.387} \quad \frac{54.20g O}{16.00g/mol} = \frac{3.387}{3.387}$$

$$= 1 \times 2 = 2 \quad = 1.496 \times 2 = 3 \quad = 1 \times 2 = 2$$



8. A 7.33-g sample of lanthanum, La, combines with oxygen to give 10.29 g of the oxide. Calculate the empirical formula of this oxide.

$$10.29g - 7.33g La = 2.96g O$$

$$\frac{7.33g La}{138.91g/mol} = .0527 \quad \frac{2.96g O}{16.00g/mol} = .1850 = \frac{.1850}{.0527} = 3.5 \times 2 = 7$$

$$= 1 \times 2 = 2$$



9. Calculate the molecular formula of a compound with the empirical formula  $CH_2O$  and a molar mass of 150 g/mol.

- [A]  $C_7H_6O_3$  [B]  $C_5H_{10}O_5$  [C]  $C_3H_6O_3$  [D]  $C_4H_8O_4$  [E]  $C_6H_{12}O_6$

$$30 \quad 150 \div 30 = 5$$

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