

Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.94 g of iron and 35.93 g of sulfur. Find the empirical formula.

$$\text{40.94 g Fe} \times \frac{1 \text{ mol Fe}}{55.845 \text{ g Fe}} = 0.733 \text{ mol Fe}$$

$$\text{35.93 g S} \times \frac{1 \text{ mol S}}{32.06 \text{ g S}} = 1.121 \text{ mol S}$$

FeS

2. A compound is found to contain 20.0 g of carbon, 26.7 g of oxygen, and 6.7 g of hydrogen. Find the empirical formula.

$$\text{20.0 g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.66 \text{ mol C}$$

$$\text{26.7 g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 1.67 \text{ mol O}$$



$$\text{6.7 g H} \times \frac{1 \text{ mol H}}{1.008 \text{ g H}} = 6.64 \text{ mol H} \quad \text{C}_1\text{O}_1\text{H}_6$$

3. A 10.0 g sample of a compound, a gas at room temperature, is found to contain 2.0 g of carbon, 2.0 g of hydrogen, and 6.0 g of oxygen. Find the empirical formula.

$$\text{2.0 g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 0.166 \text{ mol C}$$



$$\text{2.0 g H} \times \frac{1 \text{ mol H}}{1.008 \text{ g H}} = 1.98 \text{ mol H} \quad \text{C}_1\text{H}_2\text{O}_1 = \text{CH}_2\text{O}$$

$$\text{6.0 g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.375 \text{ mol O} \quad \text{C}_1\text{H}_2\text{O}_1 = \text{CH}_2\text{O}$$

4. A 20.00 g sample of an unknown compound contains 12.00 g carbon, 2.00 g hydrogen, and 6.00 g oxygen. Find the empirical formula.

$$\text{12.00 g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.00 \text{ mol C}$$



$$\text{2.00 g H} \times \frac{1 \text{ mol H}}{1.008 \text{ g H}} = 1.98 \text{ mol H}$$

$$\text{6.00 g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.375 \text{ mol O} \quad \text{C}_1\text{H}_2\text{O}_1 = \text{CH}_2\text{O}$$

Empirical Formula Pogil Answers

R Barnett



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