

Let the relative atomic mass of element Z be z .

| Element | Z | Bromine, Br |
|--------------------------|------------------|--------------------------|
| Mass (g) | 2.07 | $3.67 - 2.07 = 1.60$ |
| Number of moles of atoms | $\frac{2.07}{z}$ | $\frac{1.60}{80} = 0.02$ |

Based on the empirical formula ZBr_2 , the ratio of atoms of Z : Br is 1 : 2.

Therefore, $\frac{2.07}{z} : 0.02 = 1 : 2$

$$\frac{2.07}{z} \div 0.02 = 1 \div 2$$

$$\frac{2.07}{0.02z} = \frac{1}{2}$$

$$z = \frac{2 \times 2.07}{0.02}$$

$$= 207$$

The relative atomic mass of element Z is 207.

Empirical Formula For Copper Oxide

Jacqueline Fergusson, Richard Hart



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