Pre-lab Questions: The Determination of the Empirical Formula of a Hydrate

A hydrate is an ionic compound (ionic salt) that contains water molecules in its structure. An anhydrate (also called anhydrous salt) is the substance that remains after the water is removed from a hydrate. When a hydrate is heated, the water molecules are removed in the form of steam, leaving behind the anhydrate. To determine the formula of a hydrate experimentally, the mole ratio of the water portion to the mole ratio of the anhydrous portion must be calculated.

Questions:

During a Chem 121 experiment, a student obtained the following data after heating a sample of calcium chloride hydrate CaCl₂•xH₂O, where x refers to the number of water molecules hydrated per molecule of CaCl₂.

Data:

| Weight of crucible and lid (g) | 32.1070 |
|--|---------|
| Weight of crucible, lid and hydrated salt (g) | 36.8283 |
| Weight of crucible, lid and anhydrous salt after | |
| a) 1st heating (g) | 35,7347 |
| b) 2nd heating (g) | 35.6682 |
| c) 3rd heating (g) | 35.6679 |

Be sure to show the general formula used for each step.

1. Use this information to determine the weight of the hydrated salt (CaCl2*xH2O):

a) the weight of the anhydrous salt (CaCl₂)

Note: the weight after the 2nd and 3nd heating are close enough in value to assume that constant weight has been reached and all the water molecules have been removed

Empirical Formula Determination Lab Answers

Jicheng Xie

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