Name:	Class:	Date:

Chapter 5 Chemistry: Electrons in Atoms

Matching.

Match each item with the correct statement below.

- Pauli exclusion principle
- aufbau principle

- electron configuration
- ground state d.
- 1. each orbital has at most two electrons
- lowest energy level.
- 3. tendency of electrons to enter orbitals of lowest energy first
 - 4. arrangement of electrons around atomic nucleus

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- What is the maximum number of electrons that can occupy the n = 3 shell?
 - b.

c. 8

18

- 6. Which of the following elements is a d-block element?
 - Cd
- b.

Cs

CI

- Which element has the electron configuration Is²2s²2p⁶3d⁶⁶4s²?
 - Ag
- b. Cd.

- Circ
- 8. Which element has the following ground state electron configuration?

Co

48.8 Cu

Za

- shell is the lowest that may contain d-orbitals.
- b. 4

c. 3

- 10. In Bohr's model of the atom, where are the electrons and protons located?
 - The electrons move around the protons, which are at the center of the atom.
 - The electrons and protons move throughout the atom.
 - The electrons occupy fixed positions around the protons, which are at the center of the
 - The electrons and protons are located throughout the atom, but they are not free to move.
- 11. How many energy sublevels are in the second principal energy level?
 - h: 2

- d. 4
- 12. What is the maximum number of orbitals in the p sublevel?

c. 4 d. 5

ъ. 3

Electrons In Atoms Chapter Test A

Wolfgang Guggemos

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