

1. How did Bohr expand on Rutherford's model of the atom?

The Rutherford model of the atom was incomplete because it didn't explain how the atom's negatively charged electrons were distributed in the space surrounding the positively charged nucleus.

2. Compare the energy of an electron in the ground state and an electron in the excited state.

A valence electron in the ground state has to absorb energy in order for it to move or jump from the ground state to the excited state. This means that a ground state electron is lower energy than an excited state electron.

3. When an electron falls from a higher level to a lower level, how is the energy released?

When a valence electron absorbs energy in the form of heat or light it uses that energy to jump to an excited state (outer level). An excited (valence) electron then releases that energy in the form of light (and a little heat) and falls back to its ground state. The movement of the valence electron from the ground state to the excited state is momentary.

4. Explain how the gaseous neon atoms in a neon sign emit light.

Gaseous atom (nonmetal atoms) absorb energy in the form of high voltage electrical energy. The valence electrons of the gaseous atoms use this energy to jump from the ground state to an excited state and then emit (give off) that energy in the form of light (and a little heat).

5. List the seven colors of the visible spectrum in order of increasing energy.

Red, Orange, Yellow, Green, Blue, Indigo, and Violet

6. What is the energy difference between a photon of yellow light and a photon of violet light?

A photon of yellow light is lower energy than a photon of violet energy.

7. Determine the type of radiation (gamma rays, infrared waves, or radio waves) that has the:

- a. longest wavelength radio waves
- b. highest frequency gamma rays
- c. greatest energy gamma rays

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Victor M. Corman



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Table of Contents Electron Configuration Webquest Answer Key

1. Understanding the eBook Electron Configuration Webquest Answer Key
 - The Rise of Digital Reading Electron Configuration Webquest Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Webquest Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Webquest Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Webquest Answer Key
 - Personalized Recommendations
 - Electron Configuration Webquest Answer Key User Reviews and Ratings
 - Electron Configuration Webquest Answer Key and Bestseller Lists
5. Accessing Electron Configuration Webquest Answer Key Free and Paid eBooks

- Electron Configuration Webquest Answer Key Public Domain eBooks
- Electron Configuration Webquest Answer Key eBook Subscription Services
- Electron Configuration Webquest Answer Key Budget-Friendly Options
- 6. Navigating Electron Configuration Webquest Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Webquest Answer Key Compatibility with Devices
 - Electron Configuration Webquest Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Webquest Answer Key
 - Highlighting and Note-Taking Electron Configuration Webquest Answer Key
 - Interactive Elements Electron Configuration Webquest Answer Key
- 8. Staying Engaged with Electron Configuration Webquest Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Webquest Answer Key
- 9. Balancing eBooks and Physical Books Electron Configuration Webquest Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Webquest Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Webquest Answer Key
 - Setting Reading Goals Electron Configuration Webquest Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Webquest Answer Key
 - Fact-Checking eBook Content of Electron Configuration Webquest Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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