7&8 • Atomic Structure & Periodicity

PRACTICE TEST

$$A = 2.18 \times 10^{-18} \text{ J}$$
 $h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s}$ $h = 1.097 \times 10^{-7} \text{ m}^{-1}$ $h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s}$ $h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s}$ $h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s}$ $h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s}$ $h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s}$ $h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s}$ $h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s}$

- What wavelength corresponds to a frequency of $8.22 \times 10^9 \text{ Hz?}$
 - a) 0.307 m
- d) 0.110 m
- b) (0.0365 m 5.00ms) 27.4 m
- A radio station transmits at 110 MHz (110 x 106
 - Hz). What wavelength is this radio wave?
 - a) 3.65 x 10⁻⁵ m
- c) 3.81 x 10⁻⁵ m
- b) 3.30 m
- Which one of the following is NOT a proper unit for frequency?
 - a) Hz
 - b) 5 1
- Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen
 - a) 0.12334 m.
 - d)) 4.1029 x 10-7 m
 - b) 24.373 m
 - e) 36.559 m.
 - c) 2.7353 x 10⁻⁷ m DE= 484415197
- What is the relationship between the energy of a photon of light and its frequency?
 - mile House we

e) $E = \frac{V}{h}$

- What is the energy needed to raise an electron in the hydrogen atom from the second energy level to the third energy level? 3 - E2
 - a) 1.52 x 10⁴ J
 - b) 3.63 x 10⁻¹⁹ J en 3.03 x 10-19 J
 - c) 2.18 x 10⁻¹⁹ J -2-18 YES-15-3

 - What is the de Broglie wavelength of an electron moving at 80.0% the speed of light.

 - a) 3.03 x 10⁻¹² m c) 3.30 x 10¹¹ m

d) 4.48 x 10⁻¹⁹ J

- b) 2.42 x 10⁻¹² m d) 1.59 x 10⁻²⁵ m

 λ = 5 = 2.62 × 10⁻³ × 3.00 × 10

 8. What resultant is expected from the interference
- of the two waves shown below?







- Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?

c) 1967

dis ess.

Electron Configuration Chem Worksheet 5 6 Answers

Raffaela Di Napoli

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