

## Electron Configuration Practice

In the space below, write the standard electron configurations of the following elements:

- |    |          |  |
|----|----------|--|
| 1) | nitrogen | $1s^2 2s^2 2p^3$   |
| 2) | copper   | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^{10}$  |
| 3) | chlorine | $1s^2 2s^2 2p^6 3s^2 3p^5$   |
| 4) | calcium  | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$  |
| 5) | lead     | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$ |

In the space below, write the noble gas notation electron configurations of the following elements:

- |     |          |                             |
|-----|----------|-----------------------------|
| 6)  | barium   | $[Xe] 6s^2$                 |
| 7)  | gold     | $[Xe] 4f^{14} 5d^{10} 6s^1$ |
| 8)  | platinum | $[Xe] 4f^{14} 5d^8 6s^1$    |
| 9)  | xenon    | $[Kr] 4d^{10} 5s^2 p^6$     |
| 10) | calcium  | $[Ar] 4s^2$                 |

Determine what elements are denoted by the following electron configurations:

- |     |  |             |
|-----|--|-------------|
| 11) | $1s^2 2s^2 2p^6 3s^2 3p^3$                             | phosphorous |
| 12) | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^7$ | ruthenium   |
| 13) | $[Kr] 4d^{10} 5s^2 p^6$                                | iodine      |
| 14) | $[Xe] 4f^{14} 5d^{10} 6s^2$                            | mercury     |
| 15) | $[Rn] 7s^2$  | radium      |

Determine which of the following electron configurations are not correct:

- |     |                                 |  |
|-----|---------------------------------|--|
| 16) | $1s^2 2s^2 2p^6 3s^2 3p^5 4s^2$ | 3p <sup>5</sup> should be a 6                      |
| 17) | $1s^2 2s^2 2p^6 3s^2 2p^5$      | 1s only holds 2 and there aren't two "2p" orbitals |
| 18) | $[Cs] 5f^{14} 6d^1 7s^2$        | Cs is not a noble gas                              |
| 19) | $[Kr] 4d^{10} 5s^2 p^4$         | Correct  |
| 20) | $[Ne]$ for Neon                 | Can't use a noble gas to shortcut itself           |

# Electron Configuration And Lots More Answer Key

**Surender Kumar Sharma**

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