

## ***ELECTROCHEMISTRY***

### Basic Terminology, Construction of Electrochemical Cell and Its Representation

- Which is not true for a standard hydrogen electrode ?  
(1) The hydrogen ion concentration is 1 M  
(2) Temperature must be 25°C  
(3) Pressure of hydrogen is 1 atmosphere  
(4) All are correct
- In a galvanic cell  
(1) chemical reaction produces electrical energy  
(2) electrical energy produces chemical reaction  
(3) reduction occurs at anode  
(4) oxidation occurs at cathode
- The equation representing the process by which standard reduction potential of zinc can be defined is :  
(1)  $\text{Zn}^{2+}(\text{s}) + 2\text{e}^{-} \longrightarrow \text{Zn}(\text{s})$   
(2)  $\text{Zn}(\text{g}) \longrightarrow \text{Zn}^{2+}(\text{g}) + 2\text{e}^{-}$   
(3)  $\text{Zn}^{2+}(\text{g}) + 2\text{e}^{-} \longrightarrow \text{Zn}(\text{s})$   
(4)  $\text{Zn}^{2+}(\text{aq.}) + 2\text{e}^{-} \longrightarrow \text{Zn}(\text{s})$
- Which of the following statement is wrong about galvanic cell ?  
(1) Cathode is positive charged  
(2) Anode is negatively charged  
(3) Reduction takes place at the anode  
(4) Electrons move from anode to cathode
- KCl can be used in salt bridge as electrolyte in which of the following cells?  
(1)  $\text{Zn} \mid \text{ZnCl}_2 \parallel \text{AgNO}_3 \mid \text{Ag}$   
(2)  $\text{Pb} \mid \text{Pb}(\text{NO}_3)_2 \parallel \text{Cu}(\text{NO}_3)_2 \mid \text{Cu}$   
(3)  $\text{Cu} \mid \text{CuSO}_4 \parallel \text{AuCl}_3 \mid \text{Au}$   
(4)  $\text{Fe} \mid \text{FeSO}_4 \parallel \text{Pb}(\text{NO}_3)_2 \mid \text{Pb}$
- Which of the following has been universally accepted as a reference electrode at all temperatures and has been assigned a value of zero volt?  
(1) Standard platinum electrode  
(2) Standard copper electrode  
(3) Standard graphite electrode  
(4) Standard hydrogen electrode

# Electrochemistry Practical Manual

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*Handbook of Electrochemistry* Cynthia G. Zoski, 2006-12-11 Electrochemistry plays a key role in a broad range of research and applied areas including the exploration of new inorganic and organic compounds biochemical and biological systems corrosion energy applications involving fuel cells and solar cells and nanoscale investigations The Handbook of Electrochemistry serves as a source of electrochemical information providing details of experimental considerations representative calculations and illustrations of the possibilities available in electrochemical experimentation The book is divided into five parts Fundamentals Laboratory Practical Techniques Applications and Data The first section covers the fundamentals of electrochemistry which are essential for everyone working in the field presenting an overview of electrochemical conventions terminology fundamental equations and electrochemical cells experiments literature textbooks and specialized books Part 2 focuses on the different laboratory aspects of electrochemistry which is followed by a review of the various electrochemical techniques ranging from classical experiments to scanning electrochemical microscopy electrogenerated chemiluminescence and spectroelectrochemistry Applications of electrochemistry include electrode kinetic determinations unique aspects of metal deposition and electrochemistry in small places and at novel interfaces and these are

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