

Answer Key

PG-1: What is the coefficient of the hydroxide ion for the half reaction when the disproportionation reaction below is balanced in basic solution? $\text{BrO}^-(\text{aq}) \rightarrow \text{Br}^-(\text{aq}) + \text{BrO}_3^-(\text{aq})$ unbalanced



PG-2: What is the balanced half reaction of the species below reacting in basic solution according to the reaction below?



PG-3: What is the coefficient for water when the reaction is balanced in acidic solution? $\text{CrO}_4^{2-}(\text{aq}) + \text{Mn}^{2+}(\text{aq}) \rightarrow \text{Cr}^{3+}(\text{aq}) + \text{MnO}_4^-(\text{aq})$ unbalanced



PG-4: Which reaction occurs in the cell shown in the picture and is consistent with these electrode potentials?

The standard reduction potential of zinc is -1.26 V .
 The standard reduction potential of nickel is -0.25 V .



The cell is a galvanic cell. Electrons flow from the zinc electrode to the Nickel electrode. Reduction occurs at the Nickel electrode. Oxidation occurs at the Zinc electrode. For a reaction to be spontaneous and an electrochemical cell to be spontaneous, $\Delta G^\circ < 0$ and $E^\circ > 0$. The reaction is not spontaneous because $\Delta G^\circ > 0$ and $E^\circ < 0$. Reduction is not spontaneous and not a reduction.

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Ying-Ying Zheng



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