# PUNJAB PUBLIC SERVICE COMMISSION COMBINED COMPETITIVE EXAMINATION FOR RECRUITMENT TO THE POSTS OF PROVINCIAL MANAGEMENT SERVICE, ETC.

# CHEMISTRY (OPTIONAL) PAPER-I

TIME	ALL	OWED: 3 HOURS MAXIMUM MARKS: 100
Note:	At	tempt any FIVE questions only. Write answers in English.
Q.1:	(a)	Using a diagram, explain the variation of molecular energy with internuclear separation. (10)
0.05	(b)	What is the primary assumption of the VSEPR model? Explain the difference between square-pyramidal and trigonal bi-pyramidal arrangement in terms of VSEPR. Predict the shape of an SEr molecule.
Q.2:	(a)	What are the mineral sources for copper? Describe in detail the information which can be obtained for 3d metals using a FROST diagram. (10)
Q.85	(b)	(i) On the basis of trends in the 3d series, suggest possible M <sup>2+</sup> aqua ions for use as reducing agents, and write a balanced equation for the reaction of one of these ions with O <sub>2</sub> in acidic solution.
		(ii) Comment on the 2+ oxidation states of 4d and 5d elements. (5)
Q.3:	(a)	Explain with the help of a flow sheet diagram, the electrolytic process of caustic soda manufacturing. (10)
_==	(p)	Give a detailed comparison between dry and wet processing of cement manufacturing. (10)
Q.4:	(a)	Discuss the physical origin of quantization energy for a particle confined to moving inside a one dimensional box or on a ring. (10)
	(ь)	Use EHT to calculate the energies of the three molecular orbitals in an $H_3$ molecule ion treated as an equilateral triangle of side $R = 3.25a_0$ (170 pm). (10)
Q.5:	(a)	Calculate the work done when 50g of iron reacts with HCl to produce hydrogen gas in:  (i) a closed vessel of fixed volume, (ii) an open beaker at 25°C.
	(b)	The molar constant pressure heat capacity of a certain solid at 10K is 0.43Jk-1 mol-1. What is its molar entropy at that temperature?
Q.6:	(a)	Suggest a calibration procedure for a Na+ selective electrode, in which the values of $E_{\rm sp}$ , $\beta$ and $k_{\rm xy}$ are measured experimentally. Assume that $K^*$ is the only interfering species. (10)
	(p)	Distinguish between galvanic and electrolytic cells. (10)
Q.7:	(a)	Define Nuclear Fission and Nuclear Fusion processes.  Discuss their comparative features.
Q.853	(b)	Outline a method for preparation of CAN pellets. Distinguish between calcium superphosphate and Triple Superphosphate.

# June 2012 66ch01 Chemistry Past Paper

Jessica J Manson

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