

Chem 1020 Molarity worksheet, version 2

1. What is the molarity of a solution that contains 5.83 g LiCl in 1.50 L solution?

$$5.83 \text{ g LiCl} \times \frac{1 \text{ mol LiCl}}{42.389 \text{ g LiCl}} = 0.13754 \text{ mol LiCl}$$

$$M = \frac{\text{mol solute}}{\text{L solution}} = \frac{0.13754 \text{ mol LiCl}}{1.50 \text{ L solution}} = 0.0917 \text{ M}$$

2. What is the volume of a 0.300 M CuSO₄ solution that contains 12.4 g CuSO₄?

$$12.4 \text{ g CuSO}_4 \times \frac{1 \text{ mol CuSO}_4}{159.6 \text{ g CuSO}_4} \times \frac{1 \text{ L solution}}{0.300 \text{ mol CuSO}_4} = 0.26 \text{ L solution}$$

3. What mass of solid KBr would be needed to prepare 750.0 mL of 1.75 M KBr solution? How specifically would this solution be prepared?

$$750.0 \text{ mL KBr soln} \times \frac{1 \text{ L soln}}{1000 \text{ mL soln}} \times \frac{1.75 \text{ mol KBr}}{1 \text{ L soln}} \times \frac{119.01 \text{ g KBr}}{1 \text{ mol soln}} = 156 \text{ g KBr}$$

Prepare this solution by measuring 156 g KBr into a 750 mL volumetric flask, and adding enough water to make 750 mL solution.

4. What is the % by mass of 28.3 g of solution which contains 3.56 g NaCl?

$$\% (\text{mass}) = \frac{\text{mass solute}}{\text{mass solution}} \times 100\% = \frac{3.56 \text{ g NaCl}}{28.3 \text{ g solution}} \times 100\% = 12.6\%$$

5. What volume of methanol is required to prepare 2.0 L of a 15.0% by volume solution?

$$\% (\text{vol}) = \frac{\text{vol solute}}{\text{vol solution}} \times 100\%$$

$$\text{vol solute} = \left(\frac{\% (\text{vol})}{100\%} \right) \times \text{vol solution} = \left(\frac{15.0\%}{100\%} \right) \times 2.0 \text{ L} = 0.30 \text{ L}$$

6. What volume of 12.0 M HCl is required to make 500.0 mL of 0.400 M HCl? How specifically would this solution be prepared?

$$M_1 V_1 = M_2 V_2$$

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(0.400 \text{ M})(500.0 \text{ mL})}{(12.0 \text{ M})} = 16.7 \text{ mL}$$

Prepare this solution by measuring 16.7 mL of 12.0 M HCl into a 500 mL volumetric flask, and adding water to make 500 mL of solution.

7. If 17.0 mL of 16.0 M H₂SO₄ solution is diluted to 100.0 mL, what is the diluted concentration?

$$M_1 V_1 = M_2 V_2$$

$$M_2 = \frac{M_1 V_1}{V_2} = \frac{(16.0 \text{ M})(17.0 \text{ mL})}{(100.0 \text{ mL})} = 2.72 \text{ M}$$

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web dec 1 2014 file previews lesson plan demonstrating how to draw 3d shapes on isometric grid with modelling and examples

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web 3d number cubes for customisable size which can be used to learn about place value or to create 3d shapes slider to subdivide number bars into their factors and the to split them into individual tiles

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web students can cut and fold the nets to create 3d figures and the grid background allows for an exploration of surface area and volume concepts this product includes grid paper nets for the following five figures rectangular prism triangular prism cylinder square pyramid and cube subjects