

Half-Modern Chemistry Review

CHAPTER 11: CHEMICAL BONDING

This half-modern program concludes that level (that vast 50%) of that information for this chapter is dead.

Focus on this content, but make sure to review class notes, activities, materials, experiments, etc.

If you study this document and DISCOVERED what, you should at least be able to PLACE that item.

Now that items will be recall, examples, and/or application of this content. more

OUTCOMES

- Communicate with precision the understanding/properties learned (p2 c2)
- Communicate understanding learned to teachers and peers (p2 c2)
- EL-2 Give the Lewis dot notation, illustrate the placement bond on the drawing of molecules (p2 d, p2)
- EL-2 List important physical properties of some compounds; named molecules/compounds in some (p2 d, p2)

EL-1: INTRODUCTION TO CHEMICAL BONDING

• Introduction

- a) chemical bond → the attraction force that binds atoms or ions together
- a) ionic bonding → a force that attracts particles that have been able to attract, which demonstrates a mutual attraction, and can be
- a) covalent bonding → a bond formed when atoms share one or more pairs of valence electrons
- a) intermolecular bonding → a bond formed by attractive forces between atoms and molecules attracted to each other.
- a) polar → substance, a molecule or particle that possesses net magnetic charge, and is polarized
- a) polar covalent bond → a covalent bond in which a pair of electrons shared by two atoms is held more closely to one atom

• Chapter Highlights

- a) Most atoms are chemically bonded by ionic bonds. The most major types of chemical bonding are ionic, covalent, and metallic
- a) In general, atoms of metals, bond covalently with atoms of nonmetals, atoms of metals bond covalently with each other, and atoms of nonmetals bond covalently with each other

EL-2: COVALENT BONDING AND MOLECULAR COMPOUNDS

• Introduction

- a) molecule → a group of atoms that are held together by chemical bonds; a molecule is the smallest unit of matter that can exist by itself and retain all of its chemical properties
- a) molecular compound → a chemical compound whose simplest unit, and represents chemical formula
- a) chemical formula → a combination of chemical symbols and numbers to represent a substance
- a) molecular formula → a chemical formula that shows the number and kinds of atoms in a molecule, but not the arrangement of the atoms

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