

1. According to the Brønsted—Lowry definition, which chemical species can function *both* as an acid and as a base?

- Cl^-
- SO_4^{2-}
- NH_4^+
- HCO_3^-
- H_3O^+

2. In the reaction $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$ which is an acid—base conjugate pair?

- H_2O and HCN
- H_2O and OH^-
- CN^- and H_2O
- HCN and OH^-

3. Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX

- It is less basic than a 1 M solution of NaY.
- It is more basic than a 1 M solution of NaY.
- It yields a neutral solution.
- It is more concentrated than a 1 M solution of NaY.

4. The pH of a 0.03 M HCl solution is

a.	1.5
b.	2.5
c.	3.5
d.	12.5

5. The pH of a 1.0×10^{-3} M $\text{Ba}(\text{OH})_2$ solution at 25°C is

a.	2.7
b.	3.0
c.	11.0
d.	11.3

4. Which is the strongest acid?

- HClO
- HClO_2
- HClO_3
- HClO_4

7. Which salt reacts with water (hydrolyzes) to produce a basic solution?

a. $\text{NaC}_2\text{H}_3\text{O}_2$
b. NaNO_3
c. NH_4Cl
d. BaSO_4

8. The weakest of the bases listed is

Acid	Conjugate Base	K _a (Ionization Constant of Acid)
HCl	Cl ⁻	100% ionized
HSO ₄ ⁻	SO ₄ ²⁻	1.2 x 10 ⁻²
H ₂ S	HS ⁻	5.7 x 10 ⁻⁸
HS ⁻	S ²⁻	1.2 x 10 ⁻¹⁴

a.	Cl^-
b.	CN^-
c.	HS^-
d.	S^{2-}
e.	SO_4^{2-}

Number of hauls	<i>M. undulatus</i> (%)	<i>M. setulosus</i> (%)	<i>M. setulosus</i> + <i>M. undulatus</i> (%)
1	10	10	20
2	30	20	50
3	50	30	80
4	70	40	100
5	85	50	100
6	95	55	100
7	100	60	100
8	100	65	100
9	100	70	100
10	100	75	100

10. Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate $\text{KC}_2\text{H}_3\text{O}_2$ is less alkaline than a 0.10 M solution of potassium cyanide, KCN?

- Hydrocyanic acid is a weaker acid than acetic acid.
- Hydrocyanic acid is less soluble in water than acetic acid.
- Cyanides are less soluble than acetates.
- Acetic acid is a weaker acid than hydrocyanic acid.

11. In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH, the properties of the solution at the equivalence point will correspond exactly to the properties of

- 0.100 M sodium solution,
- 0.0500 M sodium hydroxide solution,
- 0.0500 M benzoic acid solution,
- 0.0500 M sodium benzoate solution.

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Joshua Rueda



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Tests in Print Oscar Krisen Buros, 1974

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The ... Mental Measurements Yearbook Oscar Krisen Buros,1953

Preparing for Chemistry Teaching Festo Kayima,2025-08-13 This textbook is a comprehensive chemistry didactics resource for chemistry teacher educators chemistry teachers and trainees It provides research grounded and practical based pedagogical experiences examples and frameworks for chemistry teachers as well as a foundation for planning and implementing productive chemistry lessons The book provides a conceptual and practical roadmap illuminating which didactic knowledge elements are relevant for becoming a chemistry teacher The book starts off with a pedagogically laden however experience based justification for the relevance of chemistry didactics and then progressively breaks down the different knowledge elements that form a complete set of the didactic knowledge and skill elements a teacher needs for teaching Concrete examples are provided to allow the reader to operationalize the ideas and concepts presented in the book The structure of the chapters enables the reader to engage progressively and actively with its contents and provided examples allowing a deep understanding of the diverse links between the presented topics forming a complete set of the didactic knowledge and skills relevant for successful chemistry teaching

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