

## Galvanic (Voltaic) Cells Problems

Name \_\_\_\_\_ Hr \_\_\_\_\_

1. Regarding the following reaction:



- List the species being oxidized: \_\_\_\_\_ List the species being reduced: \_\_\_\_\_
- Calculate  $E^\circ$  for this cell.
- Which species receives electrons from the anode? \_\_\_\_\_
- Which species donates electrons to the cathode? \_\_\_\_\_

2. Given the following half-reactions:



- Write the overall equation for the galvanic cell. Be sure to balance the number of electrons.
- Calculate the  $E^\circ$  for the cell.
- Designate which  $\frac{1}{2}$  reaction occurs...  
at the anode: \_\_\_\_\_ and at the cathode: \_\_\_\_\_
- Describe the direction of electron flow.

3. Find the following using a voltaic cell with the line notation:  $\text{Al}_{(\text{s})}|\text{Al}^{3+}||\text{Pb}^{2+}|\text{Pb}_{(\text{s})}$

- Calculate the  $E^\circ$  for the reaction.
- Write the balanced overall REDOX equation for the galvanic cell.

4. Using standard reduction potentials, calculate the cell potential ( $E^\circ$ ) for each of the following reactions:



# Galvanic Cell Sample Exam Questions Withe Answers

**Oswaal Editorial Board**



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