

Thermochemistry Review - Answers

1. Using the data below, calculate the amount of energy required to melt a 16.2 g aluminum pop can. Assume that the aluminum is at an initial temperature of 20°C.

- Molar heat of fusion of Al(s) = 10.7 kJ/mol
- Melting point of Al(s) = 660°C
- Heat capacity of Al(s) = 0.900 J/g°C

$$\text{mol} = 16.2 \text{ g} \div 26.98154 \text{ g/mol} = 0.600 \text{ mol}$$

$$\Delta H = (20^\circ\text{C} \rightarrow 660^\circ\text{C}) + (\text{melt})$$

$$\Delta H = (mc\Delta T) + (n\Delta H_{\text{fus}})$$

$$\Delta H = (16.2 \text{ g})(0.900 \text{ J/g}^\circ\text{C})(640^\circ\text{C}) + (0.600 \text{ mol})(10.7 \text{ kJ/mol})$$

$$\Delta H = 9331.2 \text{ J} + 6.42 \text{ kJ}$$

$$\Delta H = 9.33 \text{ kJ} + 6.42 \text{ kJ}$$

$$\Delta H = \mathbf{15.8 \text{ kJ}}$$

2. Given the thermochemical equations:



Calculate the ΔH for the reaction: $\text{CH}_3\text{OH}(\text{l}) + \frac{3}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$



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Modern Physical Organic Chemistry Eric V. Anslyn, Dennis A. Dougherty, 2006. Making explicit the connections between physical organic chemistry and critical fields such as organometallic chemistry, materials chemistry, bioorganic chemistry, and biochemistry, this book escorts the reader into an area that has been thoroughly updated in recent times.

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