

# CHAPTER 15 STUDY GUIDE

## Energy and Chemical Change

### Section 15.1 Energy

*In your textbook, read about the nature of energy.*

**In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.**

- |       |   |
|-------|---|
| _____ | 1. <i>Energy</i> is the ability to do work or produce heat.   |
| _____ | 2. The law of conservation of energy states that energy <i>can be</i> created and destroyed.  |
| _____ | 3. Chemical potential energy is energy stored in a substance because of its <i>composition</i> .  |
| _____ | 4. <i>Heat</i> is a form of energy that flows from a warmer object to a cooler object.  |
| _____ | 5. A calorie is the amount of energy required to raise the temperature of <i>one gram</i> of pure water by one degree Celsius.                          |
| _____ | 6. A <i>calorie</i> is the SI unit of heat and energy.  |
| _____ | 7. The <i>specific heat</i> of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius. |
| _____ | 8. <i>Kinetic energy</i> is energy of motion.   |
| _____ | 9. Chemicals participating in a chemical reaction contain only <i>potential energy</i> .  |
| _____ | 10. One nutritional Calorie is equal to <i>100 calories</i> .   |
| _____ | 11. One calorie equals <i>4.184 joules</i> .  |
| _____ | 12. When a fuel is burned, some of its <i>chemical potential energy</i> is lost as heat.  |
| _____ | 13. To convert kilojoules to joules, <i>divide</i> the number of kilojoules by 1000 joules/1 kilojoule.   |

**Answer the following question. Show all your work.**

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/(g·°C).

# Energy Study Guide Modeling Chemistry

**Wolfgang Guggemos**



## Energy Study Guide Modeling Chemistry:

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