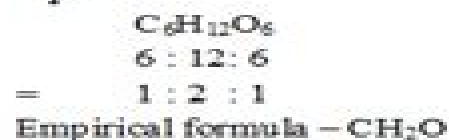
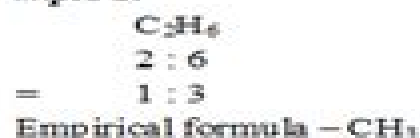


**Empirical formula:** It shows the simplest whole number ratio between the atoms of the elements in the compound.

**Example 1:**



**Example 2:**



**Example 3:**



**Practice:**

1.  $\text{C}_4\text{H}_8$
2.  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$
3.  $\text{N}_2\text{O}_4$

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**Calculate the empirical formula of the compounds having the following percentage composition:**

**Example 1:**

C-40%  
H-6.7%  
O-53.3%

**Sol:**

Element	%	% ÷ At. Wt.	Simplest ratio	simplest whole number ratio
Carbon	40	$40 \div 12 = 3.33$	$\frac{3.33}{3.33} = 1$	<b>1</b> C
Hydrogen	6.7	$6.7 \div 1$	$\frac{6.7}{3.33} = 2$	<b>2</b> H
Oxygen	53.3	$53.3 \div 16$	$\frac{3.33}{3.33} = 1$	<b>1</b> O
Empirical Formula		<b>= CH<sub>2</sub>O</b>		

# Empirical Formula Practice

**Heather Hattori, Richard H. Langley**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the authors' names.

## **Empirical Formula Practice:**

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Khomtchouk,2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields      **Chemistry** Neil D.

Jespersen,Alison Hyslop,2021-11-02 Chemistry The Molecular Nature of Matter 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text When students know what they know they are better able to learn and incorporate the material Providing a total solution through New WileyPLUS by fully integrating the enhanced etext with online assessment answer specific responses and additional practice resources The 8th edition continues to emphasize the importance of applying concepts to problem solving to achieve high level learning and increase retention of chemistry knowledge Problems are arranged in an intuitive confidence building order      *I-chemistry Iii' 2006 Ed. ,      Calculations for GCSE Chemistry* Eileen Ramsden,2002 This fully revised edition is in line with the revised 2002 National Curriculum requirements and focuses on quantitative chemistry in science Written to match all major GCSE specifications the text covers all types of numerical questions from first principles

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