

# Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.94 g of iron and 35.98 g of sulfur. Find the empirical formula.

$$\text{40.94 g Fe} \times \frac{1 \text{ mol Fe}}{55.845 \text{ g Fe}} = 0.733 \text{ mol Fe}$$

$$\text{35.98 g S} \times \frac{1 \text{ mol S}}{32.06 \text{ g S}} = 1.122 \text{ mol S}$$

FeS

2. A compound is found to contain 20.0 g of carbon, 26.7 g of hydrogen, and 66.7 g of oxygen. Find the empirical formula.

$$\text{20.0 g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.66 \text{ mol C}$$

$$\text{26.7 g H} \times \frac{1 \text{ mol H}}{1.008 \text{ g H}} = 26.5 \text{ mol H}$$



$$\text{66.7 g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 4.17 \text{ mol O} \left( \frac{4.17}{1.66} = 2.5 \right) \times 2 = 5 \text{ mol O} = 5$$

3. A 10.0 g sample of a compound contains 2.0 g of carbon, 1.0 g of hydrogen, and 7.0 g of oxygen. Find the empirical formula.

$$\text{2.0 g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 0.166 \text{ mol C}$$



$$\text{1.0 g H} \times \frac{1 \text{ mol H}}{1.008 \text{ g H}} = 0.992 \text{ mol H} \left( \frac{0.992}{0.166} = 6 \right) \times 6 = 6$$

$$\text{7.0 g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.4375 \text{ mol O} \left( \frac{0.4375}{0.166} = 26.35 \right) \times 26 = 26$$

4. A 10.00 g sample of an unknown compound contains 2.00 g carbon, 2.00 g hydrogen, and 6.00 g oxygen. Find the empirical formula.

$$\text{2.00 g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 0.166 \text{ mol C}$$



$$\text{2.00 g H} \times \frac{1 \text{ mol H}}{1.008 \text{ g H}} = 1.98 \text{ mol H}$$

$$\text{6.00 g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.375 \text{ mol O} \left( \frac{0.375}{0.166} = 22.5 \right) \times 2 = 45$$

# Empirical Formula Pogil Answers

**RC Schank**



## **Empirical Formula Pogil Answers:**

*Solving Empirical Formula Problems*, 2016 Continue your consideration of classic chemistry problems with a look at empirical formulas and examine how empirical formulas relate to molecular formulas *Empirical Formula and Ligand Formula Index for Parts 1 to 4* Leopold Gmelin, Richard Josef Meyer, Gmelin Institut für Anorganische Chemie und Grenzgebiete in der Max-Planck-Gesellschaft zur Förderung der Wissenschaften, Deutsche Chemische Gesellschaft, 1987

**Empirical Formulas** Theodore Rudolph Running, 1917 *Empirical Equations and Nomography* Dale Stroble Davis, 1943 Empirical Formulas Runni Theodore R (Theodore Rudolph), 2022-10-27 This work has been selected by scholars as being culturally important and is part of the knowledge base of civilization as we know it This work is in the public domain in the United States of America and possibly other nations Within the United States you may freely copy and distribute this work as no entity individual or corporate has a copyright on the body of the work Scholars believe and we concur that this work is important enough to be preserved reproduced and made generally available to the public We appreciate your support of the preservation process and thank you for being an important part of keeping this knowledge alive and relevant **Chemistry Equations & Answers** Speedy Publishing, 2014-08-01 One of the most important parts of learning chemistry is simply knowing all of the equations and formulas that are used in it It is important to know the exact formatting of these equations and most classes will require a student to know them for exams It is a good idea to learn these equations and formulas with the use of a study pamphlet The pamphlet can condense all of the information so a student can memorize the equations and formulas while studying **Nomographie and empirical equations**, 1952 *Empirical Formulas* Theodore Rudolph Running, 1917 **Empirical Formulas** Th.R. Running, 1917 *Chemistry Equations & Answers* Pamphlet Master, 2014-10-12 This Chemistry Equations Answers study guide is created by Pamphlet Master for students everywhere This tool has a comprehensive variety of college and graduate school topics subjects which can give you what it takes to achieve success not only in school but beyond Included in the pamphlet are Chemical Formula and Equations What is a Chemical Formula Chemical Formula and Equations Subscripts What Is A Balanced Equation How Do We Balance The Equation What About These Halves Examples of Balancing Chemical Equations *Empirical Formulas* Theodore Rudolph Running, 2006 *Empirical Equations*, 1900 **Empirical Formulas** HardPress, Running Theodore R 1866, 2013-01 Unlike some other reproductions of classic texts 1 We have not used OCR Optical Character Recognition as this leads to bad quality books with introduced typos 2 In books where there are images such as portraits maps sketches etc We have endeavoured to keep the quality of these images so they represent accurately the original artefact Although occasionally there may be certain imperfections with these old texts we feel they deserve to be made available for future generations to enjoy **Empirical Equations** University of Sydney. Department of Mechanical Engineering, 197? **A method of forming continuous empirical equations for the thermodynamic**, 1961 **New Semi-Empirical Mass**

**Formula** R. Ayres, 1961

## Decoding **Empirical Formula Pogil Answers**: Revealing the Captivating Potential of Verbal Expression

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