

Empirical Formula

Objectives

- To determine the empirical formula of a compound by use of a combination reaction.
- To practice techniques for obtaining and recording precise mass values.
- To learn safe practices of Bunsen burner use.

Introduction

The empirical formula of a compound is the simplest (smallest) whole-number ratio of moles of elements in the compound. Determination of the empirical formula requires three steps:

- Obtain the actual mass of each element in the compound
- Calculate the moles of each element from its actual mass and molar mass
- Express the ratio of moles of each element as small whole numbers

Example

Elemental analysis of a sample of an unknown crystalline solid reveals the presence of 5.52 g of sodium and 8.48 g of chlorine as the only components. The moles of each element are calculated:

$$\text{Moles of sodium: } 5.52\text{ g Na} \times \frac{1\text{ mol Na}}{22.99\text{ g Na}} = 0.240\text{ mol Na}$$

$$\text{Moles of chlorine: } 8.48\text{ g Cl} \times \frac{1\text{ mol Cl}}{35.45\text{ g Cl}} = 0.239\text{ mol Cl}$$

The mole ratio of sodium to chlorine in the compound is 0.240 : 0.239. Expressing the mole ratio as the simplest whole-number mole ratio reveals 1 mol Na : 1 mol Cl. The empirical formula is **NaCl**, and the compound is identified as *sodium chloride* (table salt).

The empirical formula of a *metal oxide* can be determined by carrying out a **combination reaction** at high temperature in air. The mass of the metal before the reaction, and the total mass of the metal oxide after the reaction are measured. From the difference between the masses, the mass of oxygen in the metal oxide is determined. The moles of each element are calculated and the mole ratio of the two elements yields the empirical formula for the metal oxide.

Example

A 5.90 g sample of titanium metal combines with oxygen in the air to form 9.84 g of titanium oxide. The empirical formula of the titanium oxide can be determined as follows:

$$\text{Moles of titanium: } 5.90\text{ g Ti} \times \frac{1\text{ mol Ti}}{47.87\text{ g Ti}} = 0.123\text{ mol Ti}$$

$$\text{Mass of oxygen: } 9.84\text{ g compound} - 5.90\text{ g Ti} = 3.94\text{ g O}$$

$$\text{Moles of oxygen: } 3.94\text{ g O} \times \frac{1\text{ mol O}}{16.00\text{ g O}} = 0.246\text{ mol O}$$

The mole ratio of titanium to oxygen in the compound is 0.123 : 0.246. Expressing the mole ratio as the simplest whole-number mole ratio reveals 1 mol Ti : 2 mol O. The empirical formula is **TiO₂**, and the compound is identified as *titanium(IV) oxide*.

Empirical Formula Of A Compound Lab Report

George W. Jr. Latimer



Empirical Formula Of A Compound Lab Report:

Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 *Laboratory Manual for Principles of General Chemistry* 11th Edition covers two semesters of a general chemistry laboratory program. The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure, selecting the proper apparatus, employing the proper techniques, systematically analyzing and interpreting the data, and minimizing inherent variables. As a result of good data, a scientific and analytical conclusion is made, which may or may not be right but is certainly consistent with the data. Experiments write textbooks; textbooks don't write experiments. A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered. Further experimentation provides additional cause-effect observations leading to an even better understanding of the experiment. The 11th edition's experiments are informative and challenging while offering a solid foundation for technique, safety, and experimental procedure. The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment. The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit. An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment. Continued emphasis on handling data is supported by the Data Analysis section.

Basic Laboratory Principles in General Chemistry Fitzgerald B. Bramwell, 1990 **Laboratory Manual** Jo A. Beran, 1990 **Bibliography of Scientific and Industrial Reports**, 1947 *Laboratory Experiments for General Chemistry* Harold R. Hunt, Toby F. Block, George M. McKelvy, 1998 Merck's Report Theodore Weicker, 1910 **Experimental Chemistry** Wilbert Hutton, 1977 Modern Experimental Chemistry George W. Jr. Latimer, 2012-12-02. Modern Experimental Chemistry provides techniques of qualitative analysis that reinforce experiments on ionic equilibria. This book includes the determination of water in hydrated salts, identification of an organic compound after determining its molecular weight, and nonaqueous titration of a salt of a weak acid. The calculation of chemical stoichiometry, calculation of thermodynamic properties by determining the change in equilibrium with temperature, and chromium chemistry are also covered. This compilation contains enough experiments for classes which have six hours of laboratory, two 3-hour meetings per week, to last two semesters. This publication is intended for chemistry students as an introductory manual to chemistry laboratory.

The Merck Report, 1910

Introduction to Chemistry, Laboratory Manual T. R. Dickson, 1994-12-23 Teaches chemistry by offering a dynamic, provocative, and relevant view of the topic and its importance to society and our daily lives. Three themes are stressed throughout the text: developing chemical thinking and a chemical vision, learning problem-solving methods, and utilizing group work and discussion activities. These themes involve and engage the students in their own learning processes; they are challenged to be active. The presentation of topics has been altered to include a new chapter which introduces the students to

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