Date

Chemistry — Unit 5 Worksheet 3 Empirical and Molecular Formulas

Show all your work when solving the following problems. Be sure to include units and label your answer.

 Find the empirical formula of a compound containing 32.0 g of bromine and 4.9 g of magnesium.

32.0 g × <u>1 mole</u> = 0.401 mole Br 79.9 g

1:2 ratio

4.9 g × 1 mole = 0.20 mole Mg 24.31 g

MgBrz

What is the empirical formula of a carbon-oxygen compound, given that a 95.2 g sample of the compound contains 40.8 g of carbon and the rest oxygen?
95.2 g - 40.8 g = 54.4 g O

 $40.8 \text{ g} \times 1 \text{ mole} = 3.40 \text{ moles C}$ 12.0 g

1:1 ratio

54.4 g x× <u>1 mole</u> = 3.40 moles O 16.0g CO

3. A compound was analyzed and found to contain 9.8 g of nitrogen, 0.70 g of hydrogen, and 33.6 g of oxygen. What is the empirical formula of the compound?

9.8 g × 1 mole = 0.70 moles N 14.0g

1:1:3 ratio

0.70 g × <u>1 mole</u> = 0.69 moles H 1.01 g HNO₃

33.6 g × 1 mole = 2.10 moles O 16.0 g

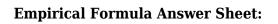
 A compound composed of hydrogen and oxygen is found to contain 0.59 g of hydrogen and 9.40 g of oxygen. The molar mass of this compound is 34.0 g/mol. Find the empirical and molecular formulas.

0.59 g × 1 mole = 0.58 mole H 1.01 g 9.40 g × 1 mole = 0.588 mole O 1:1 ratio; empirical formula: HO empirical formula mass: 17.0 g/mol molecular formula mass: 34.0 g/mol mfm is 2x efm so the molecular formula is: H-O-

16.0 g

Empirical Formula Answer Sheet

Michael Rajnik



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Table of Contents Empirical Formula Answer Sheet

- 1. Understanding the eBook Empirical Formula Answer Sheet
 - The Rise of Digital Reading Empirical Formula Answer Sheet
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Empirical Formula Answer Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - \circ Features to Look for in an Empirical Formula Answer Sheet
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Empirical Formula Answer Sheet
 - Personalized Recommendations
 - Empirical Formula Answer Sheet User Reviews and Ratings
 - Empirical Formula Answer Sheet and Bestseller Lists
- 5. Accessing Empirical Formula Answer Sheet Free and Paid eBooks
 - Empirical Formula Answer Sheet Public Domain eBooks
 - Empirical Formula Answer Sheet eBook Subscription Services

- Empirical Formula Answer Sheet Budget-Friendly Options
- 6. Navigating Empirical Formula Answer Sheet eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Empirical Formula Answer Sheet Compatibility with Devices
 - Empirical Formula Answer Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Empirical Formula Answer Sheet
 - Highlighting and Note-Taking Empirical Formula Answer Sheet
 - Interactive Elements Empirical Formula Answer Sheet
- 8. Staying Engaged with Empirical Formula Answer Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Empirical Formula Answer Sheet
- 9. Balancing eBooks and Physical Books Empirical Formula Answer Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Empirical Formula Answer Sheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Empirical Formula Answer Sheet
 - Setting Reading Goals Empirical Formula Answer Sheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Empirical Formula Answer Sheet
 - Fact-Checking eBook Content of Empirical Formula Answer Sheet
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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