

Empirical and Molecular Formulas Worksheet

Key

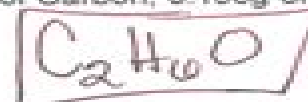
Objectives:

- be able to calculate empirical and molecular formulas

Empirical Formula

- 1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

$$\begin{array}{l} \text{C} \quad \frac{0.783\text{g C}}{12.01\text{g}} = \frac{0.065}{0.033} = 1.96 \quad 2 \\ \text{H} \quad \frac{0.196\text{g H}}{1.01\text{g}} = \frac{0.194}{0.033} = 5.88 \quad 6 \\ \text{O} \quad \frac{0.521\text{g O}}{16.00\text{g}} = \frac{0.033}{0.033} = 1 \quad 1 \end{array}$$



- 2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\text{Au} \quad \frac{89.14\text{g}}{196.97} = \frac{0.45}{0.45} = 1 \quad \text{AuO}_{1.5}$$

$$\text{O} \quad \frac{10.80\text{g}}{16\text{g}} = \frac{0.675}{0.45} = 1.5$$



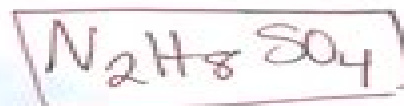
- 3) What is empirical formula if compound consists of 21.2%N, 6.1%H, 24.2%S and 48.5%O?

$$\text{N} \quad \frac{21.2\text{g}}{14.01} = 1.5 \quad 2$$

$$\text{H} \quad \frac{6.1\text{g}}{1.01\text{g}} = 6.03 \quad 8$$

$$\text{S} \quad \frac{24.2\text{g}}{32} = 0.75 \quad 1$$

$$\text{O} \quad \frac{48.5\text{g}}{16.00\text{g}} = 3.03 \quad 4$$



Molecular Formula

- 4) Empirical formula of a substance is CH_2O . Molar mass is 180. What is the molecular formula?

$$\text{EFM} = (12 + 2 \times 1 + 16) =$$

$$\frac{180}{30} = 6$$



- 5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

$$\text{C} \quad \frac{1.388}{3.585} \quad 38.7\% \rightarrow \frac{3.88\text{g}}{12.01\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{H} \quad \frac{0.345}{3.585} \quad \frac{9.62\text{g}}{1.01\text{g}} = \frac{9.52}{3.22} = 3$$

$$\text{O} \quad \frac{1.850}{3.585} \quad \frac{51.6\text{g}}{16\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{CH}_3\text{O} = \text{EFM} (12 + 3 \times 1 + 16)$$

$$\frac{62}{31.03} = 2$$



1	$\text{C}_2\text{H}_6\text{O}$
2	Au_2O_3
3	NH_4SO_4
4	$\text{C}_2\text{H}_6\text{O}_2$
5	$\text{C}_6\text{H}_{12}\text{O}_6$

Empirical Molecular Formula Answers

Arun Bahl & B S Bahl



Empirical Molecular Formula Answers:

Decoding **Empirical Molecular Formula Answers**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Empirical Molecular Formula Answers**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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