

Empirical and Molecular Formula Practice Solutions

1. A compound is found to contain 63.52 % iron and 36.48% sulfur. Find its empirical formula.

$$63.52 \text{ g Fe} \times \frac{1 \text{ mol}}{55.85 \text{ g}} = 1.137 \text{ mol Fe}$$

$$36.48 \text{ g S} \times \frac{1 \text{ mol}}{32.06 \text{ g}} = 1.138 \text{ mol S}$$

$$\frac{1.137 \text{ mol Fe}}{1.137} : \frac{1.138 \text{ mol S}}{1.137} = 1 \text{ mol Fe: } 1.001 \text{ mol S}$$

empirical formula = FeS

2. A compound is found to contain 26.56% potassium, 35.41% chromium, and the remainder oxygen. Find its empirical formula.

$$26.56 \text{ g K} \times \frac{1 \text{ mol}}{39.10 \text{ g}} = 0.6793 \text{ mol K}$$

$$36.48 \text{ g Cr} \times \frac{1 \text{ mol}}{52.00 \text{ g}} = 0.7015 \text{ mol Cr}$$

$$36.96 \text{ g O} \times \frac{1 \text{ mol}}{16.00 \text{ g}} = 2.310 \text{ mol O}$$

$$\frac{0.6793 \text{ mol K}}{0.6793} : \frac{0.7015 \text{ mol Cr}}{0.6793} : \frac{2.310 \text{ mol O}}{0.6793} = 1 \text{ mol K: } 1.033 \text{ mol Cr: } 3.401 \text{ mol O}$$

$$\rightarrow 2 \text{ mol K: } 2.066 \text{ mol Cr: } 6.802 \text{ mol O}$$

empirical formula = $\text{K}_2\text{Cr}_2\text{O}_7$

3. A 60.00 g sample of tetraethyllead, a gasoline additive, is found to contain 38.43 g lead, 17.83 g carbon, and 3.74 g hydrogen. Find its empirical formula.

$$38.43 \text{ g Pb} \times \frac{1 \text{ mol}}{207.20 \text{ g}} = 0.1855 \text{ mol Pb}$$

$$17.83 \text{ g C} \times \frac{1 \text{ mol}}{12.01 \text{ g}} = 1.485 \text{ mol C}$$

$$3.74 \text{ g H} \times \frac{1 \text{ mol}}{1.01 \text{ g}} = 3.78 \text{ mol H}$$

$$\frac{0.1855 \text{ mol Pb}}{0.1855} : \frac{1.485 \text{ mol C}}{0.1855} : \frac{3.78 \text{ mol H}}{0.1855} = 1 \text{ mol Pb: } 8.005 \text{ mol C: } 20.4 \text{ mol H}$$

empirical formula = $\text{PbC}_8\text{H}_{20}$

Empirical And Molecular Formula Calculation Practice

CH Cherryholmes



Empirical And Molecular Formula Calculation Practice:

Reviewing **Empirical And Molecular Formula Calculation Practice**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Empirical And Molecular Formula Calculation Practice**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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Table of Contents Empirical And Molecular Formula Calculation Practice

1. Understanding the eBook Empirical And Molecular Formula Calculation Practice
 - The Rise of Digital Reading Empirical And Molecular Formula Calculation Practice
 - Advantages of eBooks Over Traditional Books
2. Identifying Empirical And Molecular Formula Calculation Practice
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Empirical And Molecular Formula Calculation Practice
 - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical And Molecular Formula Calculation Practice
 - Personalized Recommendations
 - Empirical And Molecular Formula Calculation Practice User Reviews and Ratings
 - Empirical And Molecular Formula Calculation Practice and Bestseller Lists

5. Accessing Empirical And Molecular Formula Calculation Practice Free and Paid eBooks
 - Empirical And Molecular Formula Calculation Practice Public Domain eBooks
 - Empirical And Molecular Formula Calculation Practice eBook Subscription Services
 - Empirical And Molecular Formula Calculation Practice Budget-Friendly Options
6. Navigating Empirical And Molecular Formula Calculation Practice eBook Formats
 - ePub, PDF, MOBI, and More
 - Empirical And Molecular Formula Calculation Practice Compatibility with Devices
 - Empirical And Molecular Formula Calculation Practice Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Empirical And Molecular Formula Calculation Practice
 - Highlighting and Note-Taking Empirical And Molecular Formula Calculation Practice
 - Interactive Elements Empirical And Molecular Formula Calculation Practice
8. Staying Engaged with Empirical And Molecular Formula Calculation Practice
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Empirical And Molecular Formula Calculation Practice
9. Balancing eBooks and Physical Books Empirical And Molecular Formula Calculation Practice
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Empirical And Molecular Formula Calculation Practice
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Empirical And Molecular Formula Calculation Practice
 - Setting Reading Goals Empirical And Molecular Formula Calculation Practice
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Empirical And Molecular Formula Calculation Practice
 - Fact-Checking eBook Content of Empirical And Molecular Formula Calculation Practice
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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