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essentially there is no easy way to learn formulas and facts

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$$\text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$$
construct a table showing how to interpret the information in this equation in terms of a single molecule of glucose moles of reactants and products

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web there's 2 on the left and 1 on the right so we need to change the coefficient of  $\text{NH}_3$  to 2 now we have  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$  total the atoms up again on the left there is 2 N and 6 H still on the right there is 2 N and 6 H now so now all we need to do is make the left side have 6 H in total

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web chemical equations and the law of conservation of matter in the previous section the reaction between hydrogen gas and oxygen gas to produce water in the gaseous phase was shown as a chemical equation h<sub>2</sub> g o<sub>2</sub> g h<sub>2</sub> o g at the molecular level the reaction would look something like this notice that there are two oxygen atoms on the

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web oct 13 2023 this means that you will need to balance the carbon atoms first 5 use a coefficient to balance the single carbon atom add a coefficient to the single carbon atom on the right of the equation to balance it with the 3 carbon atoms on the left of the equation c<sub>3</sub> h<sub>8</sub> o<sub>2</sub> h<sub>2</sub> o<sub>3</sub> co<sub>2</sub>