- 1. Which of the following is the smallest particle of an element that can exist by itself?
 - A. Ion
 - B. Atom
 - C. Molecule
 - D. Compound
- 2. Which of the following correctly matches the subatomic particle with its charge and

location in an atom?

S 1	Subatomic Particle	Location	Charge
A.	Proton	Nucleus	Neutral
В	Neutron	Nucleus	Positive
C	Electron	Shell	Positive
D	Electron	Shell	Negative

- 3. Which of the following are responsible for bonding?
 - A. Nuclei
 - B. Protons
 - C. Neutrons
 - D. Electrons
- 4. How do you calculate the number of neutrons in an atom's nucleus?
 - A. Atomic number
 - B. Mass number atomic number
 - C. Mass number + atomic number
 - D. Number of electrons + Number of protons
- 5. Which of the following describes a cation?

L	Examples include Ca2+ and Al3+				
III.	A metal atom that has lost electrons				
IIII.	Has an equal number of protons and electrons				

- A. I and II only
- B. I and III only
- C. II and III only
- D. I. II and III
- 6. Which of the hypothetical elements shown below represents a metal?

AL.	Man
B.	ENG

- B. Di
- C. So D. Nh
- 7. Which of the following does the Bohr model represent?
 - A. A neon atom
 - B. A sodium atom
 - C. A sodium ion D. A fluorine atom
- 8. Draw Bohr diagrams for the following elements:
- a) Nitrogen atom



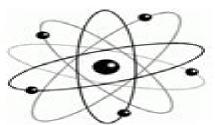
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9. Draw Lewis diagrams to show the following chemical bonds:

b) CO₂ :

a) CaCla:



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Ensheng Dong

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