

# 8 • Electron Configurations & Periodicity

## WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e <sup>-</sup> s: <u>18</u>	Element: Mg # of e <sup>-</sup> s: <u>12</u>	Element: N # of e <sup>-</sup> s: <u>7</u>	Element: Li # of e <sup>-</sup> s: <u>3</u>	Element: P # of e <sup>-</sup> s: <u>15</u>	Element: Cl # of e <sup>-</sup> s: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar  $1s^2 2s^2 2p^6 3s^2 3p^6$   
 Ar  $[\text{Ne}] 3s^2 3p^6$
- Mg  $1s^2 2s^2 2p^6 3s^2$   
 Mg  $[\text{Ne}] 3s^2$
- N  $1s^2 2s^2 2p^3$   
 N  $[\text{He}] 2s^2 2p^3$
- Li  $1s^2 2s^1$   
 Li  $[\text{He}] 2s^1$
- P  $1s^2 2s^2 2p^6 3s^2 3p^3$   
 P  $[\text{Ne}] 3s^2 3p^3$
- Cl  $1s^2 2s^2 2p^6 3s^2 3p^5$   
 Cl  $[\text{Ne}] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.

Fe  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$

Fe  $[\text{Ar}] 3d^6 4s^2$

# Electron Configuration Review Answer Sheet

**Roman Wölfel**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Environmental Chemistry Kenneth S. Overway, 2017-03-06 Covers the essentials of environmental chemistry and focuses on measurements that can be made in a typical undergraduate laboratory Provides a review of general chemistry nestled in the story of the Big Bang and the formation of the Earth Includes a primer on measurement statistics and quantitative methods to equip students to make measurements in lab Encapsulates environmental chemistry in three chapters on the atmosphere lithosphere and hydrosphere Describes many instruments and methods used to make common environmental measurements

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