

1. Which electron configuration represents an atom in an excited state?

- A) $1s^2 2s^2 2p^6 3p^1$ C) $1s^2 2s^2 2p^6 3s^2 3p^2$
B) $1s^2 2s^2 2p^6 3s^2 3p^1$ D) $1s^2 2s^2 2p^6 3s^2$

2. All of the elements in Period 3 have a total of 2 electrons in the

- A) 2s sublevel C) 2p sublevel
B) 3s sublevel D) 3p sublevel

3. Which atom in the ground state has three unpaired electrons in its outermost principal energy level?

- A) Li C) N
B) B D) Ne

4. What is the total number of valence electrons in an atom with the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^3$?

- A) 6 C) 3
B) 2 D) 5

5. Which orbital notation represents an atom of beryllium in the ground state?

- A)

1s	2s	2p		
<div style="text-align: center;">↑</div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;"> </div>
- B)

1s	2s	2p		
<div style="text-align: center;">↑↓</div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;"> </div>	<div style="text-align: center;"> </div>
- C)

1s	2s	2p		
<div style="text-align: center;">↑↓</div>	<div style="text-align: center;">↑↓</div>	<div style="text-align: center;"> </div>	<div style="text-align: center;"> </div>	<div style="text-align: center;"> </div>
- D)

1s	2s	2p		
<div style="text-align: center;">↑</div>	<div style="text-align: center;"> </div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;">↑</div>

6. Which element has a completely filled third principal energy level?

- A) Ar C) Fe
B) N D) Zn

7. Which atom in the ground state has only 3 electrons in the 3p sublevel?

- A) phosphorus C) argon
B) potassium D) aluminum

8. What is the total number of occupied principal energy levels in a neutral atom of neon in the ground state?

- A) 1 C) 3
B) 2 D) 4

9. Which sublevel configuration correctly represents a completely filled third principal energy level?

- A) $3s^2 3p^5 3d^6$ C) $3s^2 3p^6 3d^{10}$
B) $3s^2 3p^2 3d^{10}$ D) $3s^2 3p^6 3d^6$

10. Which atom in the ground state contains a partially filled 3p orbital?

- A) argon C) potassium
B) calcium D) aluminum

11. In the ground state, the atoms of elements in Period 2 all have the same number of

- A) protons C) 1s electrons
B) neutrons D) oxidation states

12. Which is the electron configuration for a neutral atom in the ground state?

- A) $1s^2 2s^2 3s^1$ C) $1s^2 2s^2 2p^6 3p^1$
B) $1s^2 2s^2 2p^4 3s^1$ D) $1s^2 2s^2 2p^6 3s^1$

13. Which represents the electron configuration of an isotope of oxygen in the ground state?

- A) $1s^2 2s^2 2p^1$ C) $1s^2 2s^2 2p^2$
B) $1s^2 2s^2 2p^2$ D) $1s^2 2s^2 2p^4$

14. The total number of sublevels in the fourth principal energy level of an atom is

- A) 1 C) 3
B) 2 D) 4

15. Which is a possible electron configuration for argon in the excited state?

- A) $1s^2 2s^2 2p^7 3s^2 3p^5$ C) $1s^2 2s^2 2p^6 3s^1 3p^2$
B) $1s^2 2s^2 2p^5 3s^2 3p^6$ D) $1s^2 2s^2 2p^6 3s^2 3p^5 4s^1$

16. Which is a correct description of the shape and spatial orientation of the p orbitals in an atom?

- A) All have the same shape but a different orientation.
B) All have the same shape and the same orientation.
C) All have a different shape and a different orientation.
D) All have a different shape but the same orientation.

17. A maximum of 6 electrons can occupy

- A) an s orbital C) a p orbital
B) an s sublevel D) a p sublevel

18. How does the ground state electron configuration of the hydrogen atom differ from that of a ground state helium atom?

- A) Hydrogen has one electron in a higher energy level.
B) Hydrogen has two electrons in a lower energy level.
C) Hydrogen contains a half-filled orbital.
D) Hydrogen contains a completely filled orbital.

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Table of Contents Electron Configuration Review Answer Key

1. Understanding the eBook Electron Configuration Review Answer Key
 - The Rise of Digital Reading Electron Configuration Review Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Review Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Review Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Review Answer Key

- Personalized Recommendations
- Electron Configuration Review Answer Key User Reviews and Ratings
- Electron Configuration Review Answer Key and Bestseller Lists
- 5. Accessing Electron Configuration Review Answer Key Free and Paid eBooks
 - Electron Configuration Review Answer Key Public Domain eBooks
 - Electron Configuration Review Answer Key eBook Subscription Services
 - Electron Configuration Review Answer Key Budget-Friendly Options
- 6. Navigating Electron Configuration Review Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Review Answer Key Compatibility with Devices
 - Electron Configuration Review Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Review Answer Key
 - Highlighting and Note-Taking Electron Configuration Review Answer Key
 - Interactive Elements Electron Configuration Review Answer Key
- 8. Staying Engaged with Electron Configuration Review Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Review Answer Key
- 9. Balancing eBooks and Physical Books Electron Configuration Review Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Review Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Review Answer Key
 - Setting Reading Goals Electron Configuration Review Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Review Answer Key

- Fact-Checking eBook Content of Electron Configuration Review Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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