

8 • Electron Configurations & Periodicity

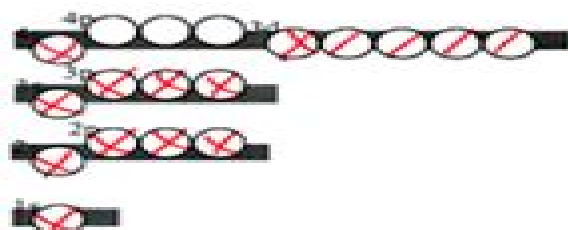
WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

| | | | | | |
|---|---|---|--|--|---|
| 1. | 2. | 3. | 4. | 5. | 6. |
| | | | | | |
| Element: Ar # of e ⁻ s: <u>18</u> | Element: Mg # of e ⁻ s: <u>12</u> | Element: N # of e ⁻ s: <u>7</u> | Element: Li # of e ⁻ s: <u>3</u> | Element: P # of e ⁻ s: <u>15</u> | Element: Cl # of e ⁻ s: <u>17</u> |

Write the electron configurations of each of these in **long form** and **short form**:

- Ar $1s^2 2s^2 2p^6 3s^2 3p^6$
 Ar $[Ne] 3s^2 3p^6$
- Mg $1s^2 2s^2 2p^6 3s^2$
 Mg $[Ne] 3s^2$
- N $1s^2 2s^2 2p^3$
 N $[He] 2s^2 2p^3$
- Li $1s^2 2s^1$
 Li $[He] 2s^1$
- P $1s^2 2s^2 2p^6 3s^2 3p^3$
 P $[Ne] 3s^2 3p^3$
- Cl $1s^2 2s^2 2p^6 3s^2 3p^5$
 Cl $[Ne] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



Electron Configuration Of Ions Practice Answer Key

Peter Atkins



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Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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