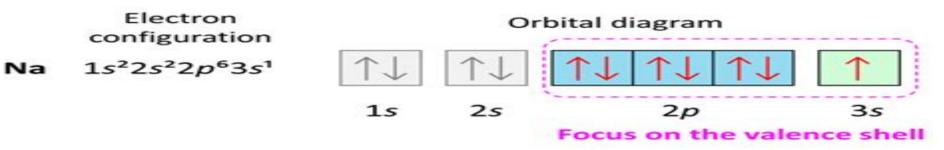
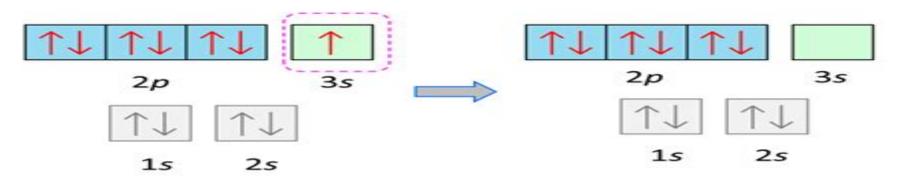
### **Electron Configuration of Main Group Cations**



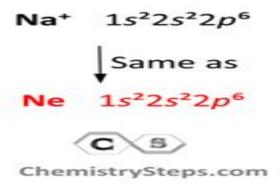
The charge is +1, therefore, one electron is lost:

The electron must be from the outermost valence orbital  $-3s^1$ 



Na  $1s^2 2s^2 2p^6 3s^1$ 

Cations of main group metals attain the electron configuration of the noble gas that precedes them in the periodic table. They are *isoelectronic* with the previous noble gas.



# **Electron Configuration Of Atoms And Ions Lab**

**RC Schank** 

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